

Kinetics: how FAST, not how much

In kinetics, we study rates of chemical reactions + sheds light on reaction mechanism

Mechanism: steps to form products

Factors affecting reaction rates:

1. **Physical state** of the reactants (the more homogenous the mixture is, the faster the reaction rate)

Ex.- liquid and liquid will react quicker than liquid and solid

In order to react, molecules must come in contact (collide) with each other (collisions/seconds)

2. **Total surface area** of the reactants (the greater the total surface area, the faster the reaction rate)

Ex.- shredded paper burns faster than a paper ball (more exposure, more time, more area)

As surface area increases, there is increased amount of contact (number of collisions) between reactants

3. **Concentration** of the reactants (as the concentration of the reactants increases, the reaction rate increases)

Ex.- 3 moles HCl > 1 mole HCl

As concentration increases, so does the likelihood that reactant molecules will collide

**For gases, concentration is directly proportional to pressure and inversely to volume of container
[concentration] is shown in brackets

4. **Temperature** of the reactants (the higher the temperature, the faster the reaction time)

At a higher temperature, there is more kinetic energy and molecules can collide more often with greater force

**Increases frequency and force of collisions (only factor with force)

5. **Presence of a Catalyst** (when you add a catalyst, it increases the reaction rate)

Catalysts create an alternate pathway and speed up reactions by changing the mechanism of the reaction. The new pathway has a lower activation energy (Eactivation)

Catalysts are NOT consumed during the course of the reaction

Ex.- $A+B \rightarrow C+A$ (A=catalyst, present before and after)

$$\frac{\Delta [Final - Initial]}{\Delta t}$$

How to determine rate of reaction: change of concentration/change in time

Example: reaction of hydrogen peroxide

Hydrogen peroxide decomposes into water and oxygen gas. The reaction is catalyzed by the enzyme catalase.



Initial concentration of H_2O_2 is 0.100 M .

$$\frac{0.100 \text{ M} - 0.050 \text{ M}}{10.0 \text{ s}} = 5.0 \times 10^{-3} \text{ M/s}$$

Calculate the rate of reaction for the reaction above. The initial concentration of H_2O_2 is 0.100 M and the final concentration is 0.050 M after 10.0 s .

The rate of reaction is the change in concentration of the reactants or products over the change in time.

Solutions Manual For Ap Chem Chemical Kinetics

Detlev Möller



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