

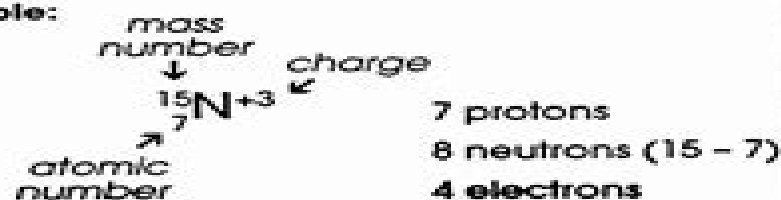
ATOMIC STRUCTURE

Name _____

An atom is made up of protons and neutrons (both found in the nucleus) and electrons (in the surrounding electron cloud). The atomic number is equal to the number of protons. The mass number is equal to the number of protons plus neutrons. In a neutral atom, the number of protons equals the number of electrons. The charge on an ion indicates an imbalance between protons and electrons. Too many electrons produces a negative charge, too few, a positive charge.

This structure can be written as part of a chemical symbol.

Example:



Complete the following chart.

Element/ Ion	Atomic Number	Atomic Mass	Mass Number	Protons	Neutrons	Electrons
H						
H ⁺						
$^{12}_{6}\text{C}$						
$^7_3\text{Li}^+$						
$^{35}_{17}\text{Cl}^-$						
$^{39}_{19}\text{K}$						
$^{24}_{12}\text{Mg}^{2+}$						
As ³⁻						
Ag						
Ag ⁺¹						
S ⁻²						
U						

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Siddappa N.Byrareddy



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Holt Chemistry Holt Rinehart & Winston,2003-01-24 **Earth Science** Samuel N. Namowitz,Nancy E. Spaulding,1989

Merrill Chemistry Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 **Exploring Earth and Space** Michael

DiSpezio,1995 A textbook exploring such aspects of matter and energy as heat electricity and nuclear chemistry with

suggested activities and review questions at the end of each chapter

The Software Encyclopedia ,1985 **Te HS&T J**

Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2004-02

Films and Other Materials for Projection Library

of Congress,1974

Current Index to Journals in Education ,1985

Library of Congress Catalogs Library of

Congress,1973

ATOMIC STRUCTURE High School Chemistry Curriculum Bud Morgan,2025-07-09 Welcome to our

comprehensive guide to Atomic Structure designed specifically for high school students studying chemistry In this book we

will explore the fascinating world of atoms the basic building blocks of matter and delve into the complexities of their

structure Chapter 1 provides a solid foundation by introducing the concept of atoms and their historical significance in the

development of modern chemistry We will examine the key scientists who contributed to our understanding of atomic

structure from Democritus to Thomson Rutherford and Bohr Moving on to Chapter 2 we will explore the structure of the

atom in more detail including the subatomic particles that make up an atom protons neutrons and electrons We will also

discuss the concept of atomic number atomic mass and isotopes In Chapter 3 we will delve into the arrangement of electrons

within an atom s electron cloud including energy levels sublevels and orbital shapes We will also discuss the principles of

electron configuration and how it impacts an atom s chemical properties Chapter 4 will focus on the periodic table and how it

organizes elements based on their atomic structure We will explore periodic trends such as atomic radius ionization energy

and electron affinity and discuss how these trends are related to an element s position on the table In our final chapter we

will discuss the significance of atomic structure in various chemical reactions including bonding stoichiometry and the

behavior of gases We will also explore the importance of understanding atomic structure in fields such as biochemistry

environmental science and materials science Throughout this book we will provide clear explanations helpful illustrations

and interactive activities to reinforce key concepts Whether you are a visual learner a hands on learner or a theoretical

learner this book is designed to meet your needs We hope that this book will not only enhance your understanding of atomic

structure but also inspire you to further explore the wonders of chemistry So let s dive in and unlock the mysteries of the

atom together

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