

Activity B:**Determining concentration**Get the Gizmo ready:

- Click **Reset**. Select **1.00 M NaOH** for the Burette and **Mystery H_2SO_4** for the Flask.
- Select **Bromthymol blue** for the Indicator.



Introduction: Adding a drop of strong acid or base into a neutralized solution is similar to adding a drop of strong acid or base to water—it causes an abrupt change in pH. By using an appropriate indicator, a chemist can tell when a solution is neutralized by monitoring its color.

Question: How is titration used to determine an unknown concentration?

1. **Measure:** Titrate the sulfuric acid analyte (H_2SO_4) with the sodium hydroxide titrant (NaOH).

How much 1.00 M NaOH is needed to neutralize the H_2SO_4 solution?

21.6 mL

2. **Interpret:** The balanced equation for the reaction of H_2SO_4 and NaOH is given at bottom right.

Based on this equation, how many moles of NaOH react with 1 mole of H_2SO_4 ?

2

3. **Manipulate:** Recall that molarity is equal to the number of moles of a substance dissolved in one liter of solution; molarity = moles ÷ volume.

- A. Write an equation for determining the number of moles of NaOH that are added to the flask based on [NaOH] and volume of NaOH titrant (mL NaOH):

Moles NaOH =

$$[NaOH] \times mL\ NaOH / 1,000\ mL$$

- B. Write a similar expression for the number of moles of H_2SO_4 in the flask based on [H_2SO_4] and the volume of H_2SO_4 (mL).

Moles H_2SO_4 =

$$[H_2SO_4] \times mL\ H_2SO_4 / 1,000\ mL$$

- C. Because there are twice as many moles of NaOH as moles of H_2SO_4 in this reaction, you can say:

$$\text{Moles NaOH} = 2 \cdot \text{Moles } H_2SO_4$$

Substitute your expressions from 3A and 3B into this equation and solve for [H_2SO_4]:

Titration Virtual Activity 1 Answer Key

Daniel F McAuley



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