



TITRATIONS 3

- 1 25.0 cm³ of 0.200 mol/dm³ sulfuric acid neutralises 18.6 cm³ of barium hydroxide solution.



- a Find the concentration of the barium hydroxide solution in mol/dm³. Give your answer to 3 significant figures.

$$\text{moles H}_2\text{SO}_4 = \text{conc} \times \text{vol (dm}^3\text{)} = 0.200 \times \frac{25.0}{1000} = 0.00500 \text{ mol}$$

$$\text{moles Ba}(\text{OH})_2 = \text{moles of H}_2\text{SO}_4 = 0.00500 \text{ mol}$$

$$\text{conc Ba}(\text{OH})_2 = \frac{\text{moles}}{\text{volume (dm}^3\text{)}} = \frac{0.00500}{\frac{18.6}{1000}} = 0.269 \text{ mol/dm}^3$$

- b Find the concentration of the barium hydroxide solution in g/dm³. Give your answer to 3 significant figures.

$$\text{conc Ba}(\text{OH})_2 = 171 \times 0.269 = 46.0 \text{ g/dm}^3$$

- c List the key steps in the practical you would carry out to find out that 18.6 cm³ of the barium hydroxide solution is needed to neutralise 25.0 cm³ of 0.200 mol/dm³ sulfuric acid.

use a pipette to place 25 cm³ of sulfuric acid
in a conical flask
add an indicator
add barium hydroxide from burette
until colour changes
go dropwise near end-point
record result
repeat

- 2 What volume of 0.0100 mol/dm³ calcium hydroxide neutralises 25.0 cm³ of 0.140 mol/dm³ sulfuric acid? Give your answer to 3 significant figures.



$$\text{moles H}_2\text{SO}_4 = \text{conc} \times \text{vol (dm}^3\text{)} = 0.140 \times \frac{25.0}{1000} = 0.00350 \text{ mol}$$

$$\text{moles Ca}(\text{OH})_2 = \text{moles H}_2\text{SO}_4 = 0.00350 \text{ mol}$$

$$\text{volume Ca}(\text{OH})_2 = \frac{\text{moles}}{\text{conc}} = \frac{0.00350}{0.0100} = 0.350 \text{ dm}^3$$

- 3 What volume of 1.50 mol/dm³ potassium hydroxide reacts with 25.0 cm³ of 1.10 mol/dm³ nitric acid? Give your answer to 3 significant figures.



$$\text{moles HNO}_3 = \text{conc} \times \text{vol (dm}^3\text{)} = 1.10 \times \frac{25.0}{1000} = 0.0275 \text{ mol}$$

$$\text{moles KOH} = \text{moles HNO}_3 = 0.0275 \text{ mol}$$

$$\text{volume KOH} = \frac{\text{moles}}{\text{conc}} = \frac{0.0275}{1.50} = 0.0183 \text{ dm}^3$$

Titration Chem Study Guide Answers

**John Krenos, Joseph Potenza, Loretta
Jones, Lynn Koplitz, Thomas Spence**



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