I. ONE REACTANT IN EXCESS - COMBUSTION

You burn an excess of butane (C₄H₁₀) in a limited amount of oxygen (O₂). If 67.2 L of CO₂ gas is formed, what volume of oxygen gas was present at the start of the reaction?

STEP 1: Balance the reaction:

STEP 2: Use a molar conversion to convert quantity of product given to moles:

STEP 3: Use a molar ratio to calculate the moles of reactant that I'm asking about:

$$3 \text{ mol CO}_2$$
 * (13 mol O_2) = 4.875 mol O_2
 (8 mol CO_2)

STEP 4: Use a molar conversion to convert moles of reactant to the units I ask for:

$$4.875 \text{ mol } O_2 \ (22.4 \text{ L} \cdot O_2) = 109.2 \text{ L} \cdot O_2 = 109 \text{ L} \cdot O_2$$

(1 mol O_2)

II. ONE REACTANT IN EXCESS - SINGLE DISPLACEMENT

14 g of lithium is combined with excess K₂SO₄. What is the mass of the new ionic compound that is formed? How do you know that this reaction will take place?

STEP I: Balance the reaction:

Reaction occurs because Li is higher up than K on the activity series.

STEP 2: Use a molar conversion to convert quantity of reactant given to moles:

STEP 3: Use a molar ratio to calculate moles of product that I'm asking about:

2 mol Li *
$$(1 \text{ mol Li}_2SO_4) = 1 \text{ mol Li}_2SO_4$$

(2 mol Li)

STEP 4: Use a molar conversion to convert moles of product to units that I ask for:

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Syed R. Qasim, Guang Zhu

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Table of Contents Simple Stoichiometry Practice Problems With Answers

- 1. Understanding the eBook Simple Stoichiometry Practice Problems With Answers
 - The Rise of Digital Reading Simple Stoichiometry Practice Problems With Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Simple Stoichiometry Practice Problems With Answers
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Simple Stoichiometry Practice Problems With Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Simple Stoichiometry Practice Problems With Answers
 - Personalized Recommendations
 - Simple Stoichiometry Practice Problems With Answers User Reviews and Ratings
 - Simple Stoichiometry Practice Problems With Answers and Bestseller Lists
- 5. Accessing Simple Stoichiometry Practice Problems With Answers Free and Paid eBooks
 - Simple Stoichiometry Practice Problems With Answers Public Domain eBooks
 - Simple Stoichiometry Practice Problems With Answers eBook Subscription Services
 - Simple Stoichiometry Practice Problems With Answers Budget-Friendly Options

- 6. Navigating Simple Stoichiometry Practice Problems With Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Simple Stoichiometry Practice Problems With Answers Compatibility with Devices
 - Simple Stoichiometry Practice Problems With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Simple Stoichiometry Practice Problems With Answers
 - Highlighting and Note-Taking Simple Stoichiometry Practice Problems With Answers
 - Interactive Elements Simple Stoichiometry Practice Problems With Answers
- 8. Staying Engaged with Simple Stoichiometry Practice Problems With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Simple Stoichiometry Practice Problems With Answers
- 9. Balancing eBooks and Physical Books Simple Stoichiometry Practice Problems With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Simple Stoichiometry Practice Problems With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Simple Stoichiometry Practice Problems With Answers
 - Setting Reading Goals Simple Stoichiometry Practice Problems With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Simple Stoichiometry Practice Problems With Answers
 - Fact-Checking eBook Content of Simple Stoichiometry Practice Problems With Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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