

EXPLORE: Visualizing Mixing with *Sugar and Salt Solutions* PhET simulation

Learning Objectives:

- Describe the process of solution formation for salts versus molecular compounds.
- Identify the differences in an ionic solution and a molecular solution

Directions: Open the *Sugar and Salt Solutions* simulation [CLICK HERE](#)

**Move the conductivity tester to the solution.
Shake some salt into the water.**

1. Does the conductivity tester light up? **Yes.**
2. What kind of substance causes the conductivity tester to light? **Electrolyte.**

**Click on the "remove salt" button.
Change the solute to sugar and shake some into the water.**

1. Did it light up? **No.**
2. What type of substance is this? **Nonelectrolyte.**

**Click on the "Micro" tab at the top of the page.
Select the sodium chloride solute.
Shake some salt into the water.**

1. What happens to the salt as it mixes with the water? **The salt ions freely move throughout the water and the salt concentration increases.**

Look at the Concentration graphic.

2. How do the concentrations of the 2 ions compare?
The salt and chloride ions equally concentrate the water.

Now move the evaporation slider to the right and hold until all the water has evaporated.

1. What happened to the concentration of the ions as the water evaporates? **The concentration of the ions increases as the water evaporates.**
2. What do you observe? **The ions also don't move as freely**
3. Is the substance, NaCl, different than when it was first added to the water? **Yes, the substance joins into one compound instead of two.**

**Click the reset button at the bottom.
Select the solute glucose.
Add the sugar to the water**

1. What do you observe? **I observe that the Glucose molecules remain intact, they don't break apart into separate atoms in the water.**
2. How does the process of solvation for sugar differ from that of salt?
The process of solvation for sugar allows the sugar molecules to stay intact in the water, while the process of solvation for salt breaks the compound apart into ions.
3. Which solution results in a greater number of particles in solution- salt or sugar? **Salt**

Use the slider at the bottom to evaporate the water.

1. What did you observe? How is the formation of sugar crystals different from the formation of salt crystals?
I observed that both the sugar and the salt particles joined together as one compound when the water was evaporated. The formation of sugar crystals consists of sugar molecules that stay

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KJ Lindholm-Leary



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