

QUESTION 4

(a) (i) Dalton's law of partial pressures states that if there is a mixture of gases which do not react chemically together, then the total pressure exerted by the mixture is the sum of the partial pressures of the individual gases that make up the mixture.

$$\begin{aligned}\text{(ii) Partial pressure of } O_2 &= \left(\frac{2}{2+3+5} \right) \times 210 \\ &= \frac{2}{10} \times 210 \\ &= 42 \text{ Nm}^{-2}\end{aligned}$$

$$\begin{aligned}\text{Partial pressure of } N_2 &= \left(\frac{5}{2+3+5} \right) \times 210 \\ &= \frac{5}{10} \times 210 \\ &= 105 \text{ Nm}^{-2}\end{aligned}$$

(b) (i) The concentration of a solution is the amount of solute in 1 dm^3 of the solution, expressed in mol/dm^3 or g/dm^3 .

$$\begin{aligned}\text{(ii) mass concentration} &= \text{molarity} \times \text{molar mass} \\ &= 0.10 \times 106 \\ &= 10.6 \text{ g/dm}^3\end{aligned}$$

(iii) Weigh 10.6 g of Na_2CO_3 accurately

- Dissolve it in a small amount of water in a volumetric flask
- Add water to make up to 1 dm^3 mark.

(c) (i) Chlorine:

- It is a greenish-yellow gas
- It is about 2.5 times denser than air

hydrogen chloride:

- Pure hydrogen chloride is a colourless gas with irritating smell
- It turns damp litmus paper red.

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