

## CHEMISTRY

2c(ii)

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

where:

 $V_1$  = Initial Volume =  $150 \text{ cm}^3$  $T_1$  = Initial Temp in kelvin =  $21^\circ\text{C} + 273 = 294\text{K}$  $T_2$  = Final Temp in kelvin =  $45^\circ\text{C} + 273 = 318\text{K}$  $V_2 = ?$ 

$$V_2 = V_1 \times \frac{T_2}{T_1}$$

$$V_2 = 150 \text{ cm}^3 \times \frac{318\text{K}}{294\text{K}}$$

$$V_2 = 150 \text{ cm}^3 \times 1.0816$$

$$V_2 = 162.24 \text{ cm}^3$$

2d.

I Sodium oxide:



II Phosphorus (III) oxide:



III Chlorine (VII) oxide



IV Magnesium oxide



# Waec 2014 Chemistry Essay And Obj

**S Marginson**



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