

Thermochemistry - Unit Test Review (Quizlet [here](#))

Chapter 6 (Thermochemistry)

1. Enthalpy

Enthalpy (H): The total energy of a system (stored in molecules and chemical bonds). A reaction's change in enthalpy is equal to the reaction's change in internal energy (U) + pressure * the reaction's change in volume (a.k.a. work) - $\Delta H_{\text{rxn}} = \Delta U_{\text{rxn}} + P\Delta V_{\text{rxn}}$. This equation can change when:

- Pressure is Constant - By subbing in the first law of thermodynamics ($\Delta U = Q + W$), we find that $\Delta H_{\text{rxn}} = Q_p$ (heat exchange @ constant pressure)
- Reaction Involves Gases - By subbing in the ideal gas law, we find that ΔH_{rxn} (ratio - ΔH_{rxn} per mol) = $\Delta U_{\text{rxn}} + \Delta nRT$. Note that $\Delta n = n_{\text{products}} - n_{\text{reactants}}$ of all gases involved in the reaction - balance the equation of the reaction and use that to find moles, and T equals temp. at SATP unless otherwise specified

2. Heat Capacity

Heat Capacity (C): The amount of heat required to raise the temperature of 1 g (specific heat capacity) OR 1 mole (molar heat capacity) of a substance by 1°C. The SHC of water is 4.18 J / g°C, and the equation relating heat capacity to heat is $Q = mC\Delta T$ (or $Q = nC\Delta T$ for MHC). When 2 substances are mixed, $Q_{\text{released}} = Q_{\text{absorbed}}$. For the substance absorbing heat, $\Delta T = T_{\text{final}} - T_{\text{initial}}$ and for the substance losing heat, $\Delta T = T_{\text{initial}} - T_{\text{final}}$

3. Calorimetry

Calorimetry: A process to determine the heat associated with a chemical reaction. There are two types of calorimetry - constant pressure and constant volume

Constant-Pressure Calorimetry: A method of calorimetry that involves reactions in the liquid phase (e.g. a solid dissolving, two liquids) at atmospheric pressure. This involves placing 2 coffee cups inside of each other, and adding the solution to the inner cups, before sealing them (with a small hole for a thermometer) on top. Any heat released is absorbed by the water, since the styrofoam acts as an

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RJ Shavelson



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