Sample Test for Unit VII

Potentially useful equations:

$$Q=mc\Delta T$$
 $Q=mL$ $Q=mH_c$ $\varepsilon_c=1-\frac{T_C}{T_H}$ $\varepsilon=\frac{W}{Q_m}$

Specific heat for water = $4186 \,\mathrm{J \, kg^{-1}C^{\circ -1}}$

1. a) Define the kilocalorie.

b) How much heat energy in kilocalories is required to change 2.3 kg of water from 38°C to 45°C?

2. a) Define Heat of Combustion.

b) A fuel-oil burning boiler is used to heat water from 20° C to 60° C. The Heat of Combustion for fuel-oil is 4.19×10^{7} J/kg. The boiler is 40% efficient. How much fuel-oil is required to heat 5000 kg of water?

- Fifty grams of glycerine at 35.0°C is mixed with 1.00 kg of water at 20.0°C. Find the specific heat of glycerine if the equilibrium temperature is measured to be 20.42°C; assume no heat losses or gains.
- 4. a) Sketch a typical phase diagram with two phase changes, labeling the axes and transitions. (Also think about what determines the slopes of the T-Q lines: How are they related to the specific heat capacity?)

b) Find the total heat energy required to change 2.0 kg of ice at -10° C into liquid water at 20°C. (Specific heat for ice = 2100J kg⁻¹C⁻¹, Latent Heat of Fusion = 3.34×10^{5} J/kg, Specific heat for water = 4186 J kg⁻¹C⁻¹, Latent Heat of Vaporization = 22.6×10^{5} J/kg).

- 5. Define and provide an example for each of the three methods of heat transfer.
- 6. a) State the First Law of Thermodynamics.
 - b) State the Second Law of Thermodynamics.
- 7. a) Draw and label a schematic diagram of a heat engine.

b) Determine the Carnot efficiency of a steam turbine engine which operates between a high temperature of 550°C and a low temperature of 90°C.

Short Answers:

a) The keal is the heat energy needed to raise 1 kg of water by 1 C° (from 14.5°C to 15.5°C).

b) $Q = mc\Delta T$, where $c_{\text{water}} = 1 \, \text{kcal} \, \text{kg}^{-1} \text{C}^{\circ -1} (= 41861 \, \text{J} \, \text{kg}^{-1} \text{C}^{\circ -1})$, so that $Q = (2.3 \, \text{kg}) (1 \, \text{kcal} \, \text{kg}^{-1} \text{C}^{\circ -1}) (45^{\circ} \text{C} - 38^{\circ} \text{C}) = 16.1 \, \text{kcal}$.

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Table of Contents Unit Vii Test V1 Answers Physics

- 1. Understanding the eBook Unit Vii Test V1 Answers Physics
 - The Rise of Digital Reading Unit Vii Test V1 Answers Physics
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Unit Vii Test V1 Answers Physics
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit Vii Test V1 Answers Physics
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Unit Vii Test V1 Answers Physics
 - Personalized Recommendations
 - Unit Vii Test V1 Answers Physics User Reviews and Ratings
 - Unit Vii Test V1 Answers Physics and Bestseller Lists
- 5. Accessing Unit Vii Test V1 Answers Physics Free and Paid eBooks
 - Unit Vii Test V1 Answers Physics Public Domain eBooks
 - Unit Vii Test V1 Answers Physics eBook Subscription Services
 - Unit Vii Test V1 Answers Physics Budget-Friendly Options
- 6. Navigating Unit Vii Test V1 Answers Physics eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit Vii Test V1 Answers Physics Compatibility with Devices
 - Unit Vii Test V1 Answers Physics Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit Vii Test V1 Answers Physics
 - Highlighting and Note-Taking Unit Vii Test V1 Answers Physics
 - Interactive Elements Unit Vii Test V1 Answers Physics
- 8. Staying Engaged with Unit Vii Test V1 Answers Physics

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Unit Vii Test V1 Answers Physics
- 9. Balancing eBooks and Physical Books Unit Vii Test V1 Answers Physics
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit Vii Test V1 Answers Physics
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit Vii Test V1 Answers Physics
 - Setting Reading Goals Unit Vii Test V1 Answers Physics
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit Vii Test V1 Answers Physics
 - Fact-Checking eBook Content of Unit Vii Test V1 Answers Physics
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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