

Study Guide – Chapter 3: Stoichiometry

Test Date: Thursday, 12/18/14

- **Chemical Equations**
 - Writing word equations
 - remember to balance & include the states of matter
 - Balancing chemical equations w/coefficients
 - 5 types of reactions
 - Combination/synthesis, Decomposition, Combustion, Single Replacement, Double Replacement
 - identify what type of reaction is happening from a given equation
 - predict the products when given the reactants or told the type of reaction...
 - $C_3H_8 + O_2 (g)$ combustion forms products of $CO_2(g)$ and $H_2O (g)$
 - Synthesis of metal + nonmetal forms an ionic compound (criss cross charges for formula)
 - Decomposition breaks compound into its elements (watch out for diatomic BrINClHOFs or P_4 / S_8 solids)
- **Formula Weight (unit: amu) = Molar Mass (unit: g/mol)**
 - Calculations of formula weights and/or molar masses
 - % by mass = % composition calculations
- **Determining Empirical & Molecular Formulas**
 - Know the difference between an empirical and a molecular formula
 - Three types of problems
 - Normal problems will give % composition data or mass data to start
 - Combustion Analysis (like the quiz)
 - Hydrate Determination
 - Use the rhyme to help you remember the steps!
- **Stoichiometric Calculations**
 - Dimensional Analysis!!!
 - Avogadro's number (use for converting to or from # of atoms/molecules) → after molecules you can go to specific atom within the compound by looking at the chemical formula (ex: 4 H atoms per 1 molec. CH_4)
 - use the molar mass to go from mass to moles and/or moles to mass
 - Molar Ratios: use coefficients from a balanced chemical equation to go between any two different substances in a reaction (section 3.6)
- **Limiting Reactant Problems**
 - through calculations, identify which is the limiting reactant and which is the excess reactant
 - calculate the amount of product formed by the limiting reactant (this is the maximum or theoretical yield but you ALWAYS choose the LESSER quantity calculated)
 - determine how much of the excess reactant is used and how much is left over after the reaction is completed
 - calculate % yield
 - calculate theoretical yield and given the actual yield (from lab) of the substance
 - given % yield, you can determine actual yield collected in the lab

Unit 11 Stoichiometry Study Guide Answer Key

Katrin Zwirgmaier



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Table of Contents Unit 11 Stoichiometry Study Guide Answer Key

1. Understanding the eBook Unit 11 Stoichiometry Study Guide Answer Key
 - The Rise of Digital Reading Unit 11 Stoichiometry Study Guide Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 11 Stoichiometry Study Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 11 Stoichiometry Study Guide Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 11 Stoichiometry Study Guide Answer Key
 - Personalized Recommendations
 - Unit 11 Stoichiometry Study Guide Answer Key User Reviews and Ratings
 - Unit 11 Stoichiometry Study Guide Answer Key and Bestseller Lists
5. Accessing Unit 11 Stoichiometry Study Guide Answer Key Free and Paid eBooks
 - Unit 11 Stoichiometry Study Guide Answer Key Public Domain eBooks
 - Unit 11 Stoichiometry Study Guide Answer Key eBook Subscription Services
 - Unit 11 Stoichiometry Study Guide Answer Key Budget-Friendly Options

6. Navigating Unit 11 Stoichiometry Study Guide Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 11 Stoichiometry Study Guide Answer Key Compatibility with Devices
 - Unit 11 Stoichiometry Study Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 11 Stoichiometry Study Guide Answer Key
 - Highlighting and Note-Taking Unit 11 Stoichiometry Study Guide Answer Key
 - Interactive Elements Unit 11 Stoichiometry Study Guide Answer Key
8. Staying Engaged with Unit 11 Stoichiometry Study Guide Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 11 Stoichiometry Study Guide Answer Key
9. Balancing eBooks and Physical Books Unit 11 Stoichiometry Study Guide Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 11 Stoichiometry Study Guide Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Unit 11 Stoichiometry Study Guide Answer Key
 - Setting Reading Goals Unit 11 Stoichiometry Study Guide Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unit 11 Stoichiometry Study Guide Answer Key
 - Fact-Checking eBook Content of Unit 11 Stoichiometry Study Guide Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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