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# Valence Electrons

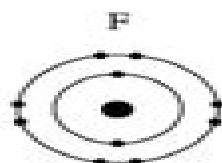
1 How many electrons are present in the following elements?

- a Carbon \_\_\_\_ b Chlorine \_\_\_\_ c Indium \_\_\_\_ d Calcium \_\_\_\_  
e Tungsten \_\_\_\_ f Antimony \_\_\_\_ g Titanium \_\_\_\_ h Tin \_\_\_\_

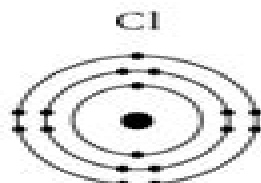
2 Using the periodic table, determine the number of valance electrons for the elements listed below.

- a Copper \_\_\_\_ b Xenon \_\_\_\_ c Hydrogen \_\_\_\_ d Aluminum \_\_\_\_  
e Barium \_\_\_\_ f Iodine \_\_\_\_ g Carbon \_\_\_\_ h Phosphorous \_\_\_\_

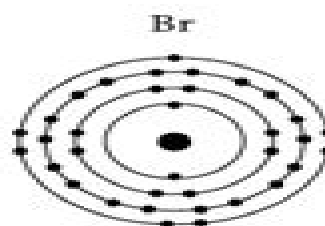
3 Circle the valence electrons only. What is the atomic number?



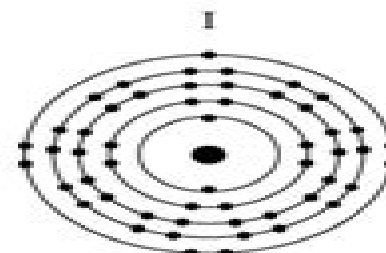
Atomic number(Z) =



Atomic number(Z) =



Atomic number(Z) =



Atomic number(Z) =

4 Fill in the table with correct values.

Element	Ion symbol	Number of Protons	Number of electrons	Charge	Ion type
Fluorine	F <sup>-</sup>			-1	Anion
		53	54		
		16		-2	
Potassium				+1	
Calcium					

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**Kaplan Test Prep**



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