

EASEL BONDING QUIZ

Multiple Choice: Select the most correct answer. (15 K)

1. An atom of argon rarely bonds to an atom of another element because an argon atom has:
A) 8 valence electrons B) 2 electrons in the 1st level
C) 3 energy levels D) 22 neutrons
2. The concept of multiple covalent bonds is used to explain the molecular formula of:
A) O₂ B) NaCl C) C₂H₄ D) F₂ E) NaOH
3. What type of bond is likely to form between the elements molybdenum, Mo and nitrogen, N based on electronegativities?
A) Non-polar covalent B) Polar covalent C) Ionic D) None of the above
4. In which pair do both compounds exhibit ionic bonding?
A) SO₂, HCl B) KNO₃, CH₄ C) NaF, K₂SO₄ D) KCl, CO₂ E) NaCl, H₂O
5. What type of bonding occurs between phosphorus and hydrogen within a molecule of PH₃?
A) ionic B) covalent C) hydrogen bonding D) metallic
6. What occurs in a triple covalent bond?
A) Two atoms share two electrons.
B) Two atoms share two pairs of electrons.
C) Two atoms share three pairs of electrons.
D) Two atoms share three electrons
7. A molecule with a single covalent bond is _____.
A) CO₂ B) CO C) Cl₂ D) N₂
8. A solid has a very high melting point, is brittle, and has poor electrical conduction. This is a(n) _____ solid.
A) ionic B) molecular C) metallic
D) molecular and ionic E) metallic and ionic
9. Which of the following statements contrasting covalent bonds and ionic bonds is correct?
A) Covalent bonds usually involve 2 nonmetals and ionic bonds usually involve 2 metals.
B) Covalent bonds usually involve 2 metals and ionic bonds usually involve a metal and a nonmetal.
C) Covalent bonds usually involve a metal and a nonmetal and ionic bonds involve 2 nonmetals.
D) Covalent bonds usually involve 2 nonmetals and ionic bonds usually involve a metal and a nonmetal.

Unit 4 Chemical Bonding Name Quiz Answer Key

Annelies Wilder-Smith



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