

Name: KEY

Valence Electrons Practice

The **valence electrons** are the electrons in the outermost energy level. They are always the outermost electrons.

Atoms that LOSE electrons are called **cations** and have an overall **positive** charge.

Atoms that GAIN electrons are called **anions** and have an overall **negative** charge.



Determine the element's number of valence electrons (# of electrons in the "outermost" shell).

Example: Carbon has 4 valence electrons, carbon 4

1. fluorine 7

2. lithium 1

3. phosphorus 5

4. francium 1

5. calcium 2

6. carbon 4

7. nitrogen 5

8. iodine 7

9. silicon 4

10. oxygen 6

11. argon 8

12. barium 2

13. potassium 1

14. aluminum 3

15. helium 2

16. hydrogen 1

17. magnesium 2

18. xenon 8

19. sulfur 6

20. boron 3

21. bromine 7

↑_{2,3}
↓
*not part of Unit 2 exam

22. What are valence electrons of an atom used for?

Determining how elements bond together

23. Which groups on the periodic table want to lose electrons? What kind(s) of elements are these groups? Group 1, 2 elements, Alkali metals and Alkali earth metals

24. Which groups on the periodic table want to gain electrons? What kind(s) of elements are these groups? Groups 5, 6, 7, non-metals

25. Why do the elements from question 23 want to lose electrons? Why do the elements from question 24 want to gain electrons? (HINT: The answer is the same for both questions.)

By losing 1 or 2 electrons these groups form more stable ion having the electron configuration of noble gases

26. How many valence electrons do the Alkaline Earth metals have? Are these elements very reactive? Explain your answer. 2, yes

27. What is the name of the atoms that lose electrons? What is their overall charge?

Cation (+) charge

28. What is the name of the atoms that gain electrons? What is their overall charge?

Anions (-) charge

Valence Electrons Assessment Key

AW Rasmussen



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A Practical Guide to Understanding, Managing, and Reviewing Environmental Risk Assessment Reports Sally L. Benjamin, David A. Belluck, 2001-02-21 A Practical Guide to Understanding Managing and Reviewing Environmental Risk Assessment Reports provides team leaders and team members with a strategy for developing the elements of risk assessment into a readable and beneficial report The authors believe that successful management of the risk assessment team is a key factor is quality report

Chemistry John Olmsted, Greg Williams, Robert C. Burk, 2020 Chemistry 4th Edition is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers and distinguish this text from other offerings It more accurately reflects the curriculum of most Canadian institutions Chemistry is sufficiently rigorous while engaging and retaining student interest through its accessible language and clear problem solving program without an excess of material and redundancy

Applied Electronics John Morris, 1996-11-29 This book provides a sound introduction to basic electronic concepts in a lively and practical format It effectively meets the needs of both the electronics option of the advanced GNVQ in engineering and the BTEC National certificate in electronics and includes hands on practical investigations and self test questions which will appeal to a wide range of readers Applied Electronics employs user friendly text and a non mathematical approach to develop the reader's ability and understanding of the principles of analogue and digital electronics Beginning with the semiconductor devices themselves it progresses through amplifiers and power supplies to combinational and sequential logic

Chemistry John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant

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Industrial Information and Design Issues Jacques-Emile Dubois, Nahum Gershon, 2012-12-06 J E DUBOIS and N GERSHON The first volume of this series The Information Revolution Impact on Science and Technology emphasized the importance of data sharing and fast communication and the advantages l f current hypertext developments in creating new and flexible data access Volume II Modeling Complex Data for Creating Information dealt in particular with the specific constraints of science and technology data including imprecision and uncertainty It also provided representation and handling tools and object oriented programming technology for developing data systems The papers presented in this third volume are concerned with the very specific information problems of the technical and competitive industrial world Here production and selling rely on creative design information processing special up to date data search knowledge comprehension and fast action all essential for decision making steps The following topics are discussed in this volume Cognition and Recognition in Design Knowledge Based Systems KBS Evaluation Modeling Tools for Knowledge Discovery Standards and CAD Computer Aided Design Aspects of Industrial Exchange and Specifications Information Seeking Strategies of Selective Access to Intelligent Information Special Information Resources Complex Databases Most of these topics inspired by the symposium on Communication and Computer Aided Systems held during the 14th International CODATA Conference deal with systemic components used by various up to date industries in development strategies

Physical Inorganic Chemistry S. F. A. Kettle, 2013-11-11 GEORGE CHRISTOU Indiana University Bloomington I am no doubt representative of a large number of current inorganic chemists in having obtained my undergraduate and postgraduate degrees in the 1970s It was during this period that I began my continuing love affair with this subject and the fact that it happened while I was a student in an organic laboratory is beside the point I was always enchanted by the more physical aspects of inorganic chemistry while being captivated from an early stage by the synthetic side and the measure of creation

with a small c that it entails I nevertheless found the application of various theoretical spectroscopic and physicochemical techniques to inorganic compounds to be fascinating stimulating educational and downright exciting The various bonding theories for example and their use to explain or interpret spectroscopic observations were more or less universally accepted as belonging within the realm of inorganic chemistry and textbooks of the day had whole sections on bonding theories magnetism kinetics electron transfer mechanisms and so on However things changed and subsequent inorganic chemistry teaching texts tended to emphasize the more synthetic and descriptive side of the field There are a number of reasons for this and they no doubt include the rise of diamagnetic organometallic chemistry as the dominant subdiscipline within inorganic chemistry and its relative narrowness vis d vis physical methods required for its prosecution

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Table of Contents Valence Electrons Assessment Key

1. Understanding the eBook Valence Electrons Assessment Key
 - The Rise of Digital Reading Valence Electrons Assessment Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Valence Electrons Assessment Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Valence Electrons Assessment Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Valence Electrons Assessment Key
 - Personalized Recommendations
 - Valence Electrons Assessment Key User Reviews and Ratings
 - Valence Electrons Assessment Key and Bestseller Lists
5. Accessing Valence Electrons Assessment Key Free and Paid eBooks
 - Valence Electrons Assessment Key Public Domain eBooks
 - Valence Electrons Assessment Key eBook Subscription Services
 - Valence Electrons Assessment Key Budget-Friendly Options

6. Navigating Valence Electrons Assessment Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Valence Electrons Assessment Key Compatibility with Devices
 - Valence Electrons Assessment Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Valence Electrons Assessment Key
 - Highlighting and Note-Taking Valence Electrons Assessment Key
 - Interactive Elements Valence Electrons Assessment Key
8. Staying Engaged with Valence Electrons Assessment Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Valence Electrons Assessment Key
9. Balancing eBooks and Physical Books Valence Electrons Assessment Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Valence Electrons Assessment Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Valence Electrons Assessment Key
 - Setting Reading Goals Valence Electrons Assessment Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Valence Electrons Assessment Key
 - Fact-Checking eBook Content of Valence Electrons Assessment Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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