



- a. How much heat is released during the formation of 9.6 moles of CO_2 ?

$$9.6 \text{ mol CO}_2 \times \frac{5638 \text{ kJ}}{12 \text{ mol CO}_2} = \text{Answer } \underline{4510.4 \text{ kJ}}$$

- b. How much heat is released during the formation of 0.036 moles of H_2O ?

$$0.036 \text{ mol H}_2\text{O} \times \frac{5638 \text{ kJ}}{11 \text{ mol H}_2\text{O}} = 18.45 \text{ kJ}$$

Answer 18.45 kJ

- c. If 1026 grams of $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ are consumed, how much heat is released?

$$1026 \text{ g C}_{12}\text{H}_{22}\text{O}_{11} \times \frac{1 \text{ mol}}{342.0 \text{ g}} = 3 \text{ mol C}_{12}\text{H}_{22}\text{O}_{11} \times \frac{5638 \text{ kJ}}{1 \text{ mol C}_{12}\text{H}_{22}\text{O}_{11}} = \text{Answer } \underline{16914 \text{ kJ}}$$

- d. If 23.76 grams of CO_2 are produced, how much heat is released?

$$23.76 \text{ g CO}_2 \times \frac{1 \text{ mol CO}_2}{44.0 \text{ g CO}_2} = 0.54 \text{ mol CO}_2 \times \frac{5638 \text{ kJ}}{12 \text{ mol CO}_2} = \text{Answer } \underline{253.71 \text{ kJ}}$$

7. Calculate the amount of heat (in Joules) required to warm 350.0 g of water from 30°C to 35°C . (Heat Capacity (C) for H_2O is $4180 \text{ J/kg} \cdot ^\circ\text{C}$)

$$\text{Heat} = m \cdot C \cdot \Delta t$$

$$= 0.350 \text{ kg} \times 4180 \frac{\text{J}}{\text{kg} \cdot ^\circ\text{C}} \times 5^\circ\text{C} = \text{Answer } \underline{7315 \text{ J}}$$

8. 35.112 kJ of heat are added to a 500.0 gram sample of water initially at 7°C . Calculate the final temperature of the water sample. Be careful with units!

$$\text{Heat} = m \cdot C \cdot \Delta t$$

$$35,112 \text{ J} = 0.500 \text{ kg} \times 4180 \frac{\text{J}}{\text{kg} \cdot ^\circ\text{C}} \times \Delta t^\circ\text{C}$$

$t_{\text{final}} = t_{\text{init}} + \Delta t$

Answer $t_{\text{final}} = 7 + 16.8 = \underline{23.8^\circ\text{C}}$

$$\Delta t = \frac{35112}{(0.500 \times 4180)} = \underline{16.8^\circ\text{C}}$$

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Power Practice: Science, Gr. 3-4, eBook Marilyn Marks, 2005-02-01

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Javier García-Martínez, Elena Serrano-Torregrosa, 2015-05-04

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web we then changed the volume between 5ml and 20ml and recorded the pressure at each this part of the lab showed that volume and pressure are inversely proportionate as the pressure would decrease if the volume was increased part b of the

experiment looked at the relationship between temperature and pressure

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web he concluded that at a constant pressure the volume of the gas would expand at a particular rate for each degree of temperature rise that being 1 267 per degree in 1874 victor regnault obtained better experimental results showing this number to be 1 273 and concluded that the pressure would approach zero at 1 273 degrees c

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web the volume of 1 00 mol 1 00 mol of any gas at stp standard temperature 273 15 k and pressure 1 atm is measured to be 22 414 l 22 414 l we can substitute 101 325 kpa 101 325 kpa for pressure 22 414 l 22 414 l for volume and 273 15 k 273 15 k for temperature into the ideal gas equation and solve for r r

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web chemists sometimes make comparisons against a standard temperature and pressure stp for reporting properties of gases 273 15 k and 1 atm 101 325 kpa 1 at stp one mole of an ideal gas has a volume of about 22 4 l this is referred to as the standard molar volume figure 9 18

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used and with it we can explore it by seeing how the pressure and volumes changed in the tubes each time that we

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web volume pressure kpa linear pressure kpa volume ml pressure kpa in part b we observed the relationship between temperature and pressure the pressure of the 125 ml flask was at 101 36 kpa when the temperature was at 23 2 c

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web pressure increases even though there were errors in the experiment the goal of the experiment was met the goal of declaring what type of relationship it is was clear using the data pressure kpa temperature c temperature k constant k p t 95 82 kpa 16 0 c 289 k 0 33 97 90 kpa 19 0 c 292 k 0 34 98 97 kpa 99 22 kpa 101 4 kpa 19 1 c

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web figure pageindex 5 the empirically determined relationships among pressure volume temperature and amount of a gas the thermometer and pressure gauge indicate the temperature and the pressure qualitatively the level in the flask indicates the volume and the number of particles in each flask indicates relative amounts

