

Concentration Practice Problems

% Solution

1. What is the concentration (% solution) of a solution that contains 0.78 grams of MgCl_2 dissolved in 10. grams of water?

$$\% = \frac{\text{mass of solute}}{\text{mass of sol'n}} = \frac{0.78 \text{ g}}{10.78 \text{ g}} = 7.24\%$$

2. How many grams of solid are needed to prepare a 10.0% NaCl solution in 200. grams of water?

$$0.10 = \frac{x}{x+200}$$

$$0.10(x+200) = x$$

$$\begin{array}{r} 0.10x + 20 = x \\ -0.10x \quad \quad -0.10x \\ \hline 20 = 0.9x \end{array}$$

$$\frac{20}{0.9} = \frac{0.9x}{0.9}$$

$$x = 22.2 \text{ g NaCl}$$

3. What is the % solution of a mixture that contains 0.0500 moles of NaI dissolved in 25.0 g of water?

$$\frac{0.0500 \text{ moles NaI}}{1} \times \frac{150.0 \text{ g}}{1 \text{ mole}} = 7.5 \text{ g NaI}$$

$$\frac{7.5 \text{ g NaI}}{(7.5 + 25) \text{ g Sol'n}} = 23.1\%$$

Molarity

4. What is the molarity of a solution that contains 85.0 g of iodine (I_2) dissolved in 250. mL of ethanol solution?

$$\frac{85.0 \text{ g}}{1} \times \frac{1 \text{ mole}}{254 \text{ g}} = 0.335 \text{ mole}$$

$$M = \frac{0.335 \text{ moles}}{0.250 \text{ L}} = 1.34 \text{ M}$$

5. How many grams of $\text{Ca}(\text{NO}_3)_2$ are need to produce 1500. mL of a 0.75 M solution?

$$M = \frac{\text{mol}}{\text{L}}$$

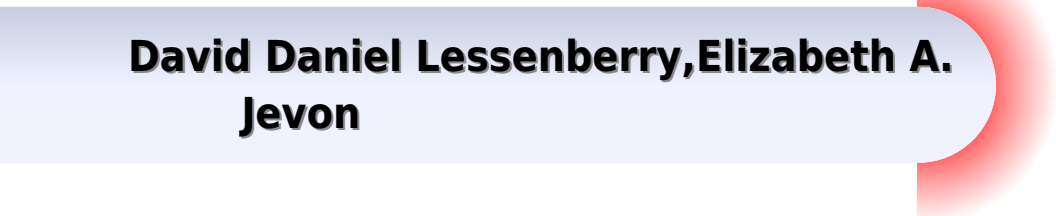
$$0.75 = \frac{x}{1.500 \text{ L}}$$

$$x = 1.125 \text{ mol}$$

$$\frac{1.125 \text{ mol}}{1} \times \frac{164.1 \text{ g}}{1 \text{ mol}} = 184.6 \text{ g Ca}(\text{NO}_3)_2$$

Skill Practice 39 Concentration Practice Answers

**David Daniel Lessenberry, Elizabeth A.
Jevon**



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Table of Contents Skill Practice 39 Concentration Practice Answers

1. Understanding the eBook Skill Practice 39 Concentration Practice Answers
 - The Rise of Digital Reading Skill Practice 39 Concentration Practice Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Skill Practice 39 Concentration Practice Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skill Practice 39 Concentration Practice Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Skill Practice 39 Concentration Practice Answers
 - Personalized Recommendations

- Skill Practice 39 Concentration Practice Answers User Reviews and Ratings
- Skill Practice 39 Concentration Practice Answers and Bestseller Lists
- 5. Accessing Skill Practice 39 Concentration Practice Answers Free and Paid eBooks
 - Skill Practice 39 Concentration Practice Answers Public Domain eBooks
 - Skill Practice 39 Concentration Practice Answers eBook Subscription Services
 - Skill Practice 39 Concentration Practice Answers Budget-Friendly Options
- 6. Navigating Skill Practice 39 Concentration Practice Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Skill Practice 39 Concentration Practice Answers Compatibility with Devices
 - Skill Practice 39 Concentration Practice Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Skill Practice 39 Concentration Practice Answers
 - Highlighting and Note-Taking Skill Practice 39 Concentration Practice Answers
 - Interactive Elements Skill Practice 39 Concentration Practice Answers
- 8. Staying Engaged with Skill Practice 39 Concentration Practice Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Skill Practice 39 Concentration Practice Answers
- 9. Balancing eBooks and Physical Books Skill Practice 39 Concentration Practice Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skill Practice 39 Concentration Practice Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Skill Practice 39 Concentration Practice Answers
 - Setting Reading Goals Skill Practice 39 Concentration Practice Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Skill Practice 39 Concentration Practice Answers
 - Fact-Checking eBook Content of Skill Practice 39 Concentration Practice Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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