Dr. - Heder Mass

- 70. Come Ser the following:
 - a. If the formula of a compound is known (e.g. CytyCl), and you find the masses of the component atoms from the Periodic Table IC = 12 011g, H = 1 0079, & Cl = 35.455g). HOW would you find the molar mass of the compound?
 - b. Calculate the molar mass of one mole of CultuCl.
 - c. Calculate the mass of 5 miles of CJHJCI

E - Percent Composition of Compounds

- Provide the word equation describing how you solve the mass percent for a given element in a compound.
- 9) Penicillin is an important antibiotic that has the chemical formula C_wH_wN_ySO_w. Calculate the mass
- percent of EACH element in this compound.

F - Empirical Formulas

- 10) Consider the following:
- Taking in to account the formula for glucose, C_aH_{ay}O_a, what information can be gethered about this compound?
 - b. What ratio of CH O is present when considering the formula above?
 - c. Now take in to account the following Formulas: CuHuOu, CuHuOu, and CuHuOu. What do these formulas have in common with glucose?
- 11) Using the word "ratio" in the descriptions, explain what the difference is between empirical and

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and H_2O Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet

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