

b) If the rate of consumption of magnesium is 5.0 x 10⁻⁹ mol/s, find the rate of consumption of HCl in moles/s.

c) If the rate of consumption of magnesium is 5.0 x 10⁻⁹ mol/s, find the rate of production of H₂ in g/s.

d) If the rate of consumption of magnesium is 5.0 x 10⁻⁹ mol/s, find the rate of production of H₂ in L/s (@STP).

e) If the rate of consumption of magnesium is 5.0 x 10⁻⁹ mol/s, find the mass of Mg consumed in 5.0 minutes.

- When butane (C₄H₁₀) is burned in air (axygen), the products carbon diaxide and water are formed.
 - a) Write a balanced formula equation for this reaction.

b) If butane is consumed at an average rate of 0.116 grams/s, determine the rate of production of CO₂ in g/s.

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Wenbin Ji

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