Chapter Review

USING KEY TERMS

For each pair of terms, explain how the meanings of the terms differ.

L solid and liquid

Solid is the state of matter in which the substance has a definite shape and vol. Liquid is the state in which the substance takes the shape of the

2. Boyle's law and Charles's law

Boyles law states that when pressure of a gas increases(at a constant temp), its volume decreases. Charles law says when temp of gas increas, the volume

3. evaporation and boiling

Evaporating us the change from a liquid to a gas at the surface of the liquid, boiling is the change of state throughout the liquid.

4. condensation and sublimation

Condensation, is the change from a g to I, and sublimation is change from a s- g.

UNDERSTANDING KEY IDEAS

Multiple Choice

- 5. Which of the following statements best describes the particles of a liquid?
 - The particles are far apart and moving fast.
 - b. The particles are close together but moving past each other.
 - c. The particles are far apart and moving slowly.
 - d. The particles are closely packed and vibrating in place.
- 6. Which of the following statements describes what happens as the temperature of a gas in a balloon increases?
 - The speed of the particles decreases.
 - The volume of the gas increases, and the speed of the particles increases.
 - c. The volume of the gas decreases.
 - d. The pressure of the gas decreases.
- 7. Boiling points and freezing points are examples of
 - a. chemical properties.

c. energy.

b. physical properties.

d. matter.

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Carol L. McClelland

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