BOYLE'S LAW

Name _____

Boyle's Law states that the volume of a gas varies inversely with its pressure if temperature is held constant. (If one goes up, the other goes down.) We use the formula:

 $P_1 \times V_1 = P_2 \times V_2$

Solve the following problems (assuming constant temperature). ,...

- 25L .4'/atm

 1. A sample of oxygen gas occupies a volume of 260. mL at 740. terr pressure. What volume will it occupy at 800. terr pressure?
- A sample of carbon dioxide occupies a volume of 3.50 liters at 125 kPa pressure. What pressure would the gas exert if the volume was decreased to 2.00 liters?
- A 2.0 liter container of nitrogen had a pressure of 3.2 atm. What volume would be necessary to decrease the pressure to 1.0 atm?
- Ammonia gas occupies a volume of 460. mt. at a pressure of 220. mm Hg. What volume will it occupy at standard pressure?
- 5. A 175 mL sample of neon had its pressure changed from 75 kPa to 150 kPa 1.5 atm.
 What is its new volume?
- 6. A sample of hydrogen at 1.5 atm had its pressure decreased to 0.50 atm producing a new volume of 759 ml. What was its original volume?
- Chlorine gas occupies a volume of 1.2 liters at 720 terr pressure. What volume will it occupy at 1 atm pressure?
- Fluorine gas exerts a pressure of \$\frac{\frac{\chi_0}{\chi_0}\text{Text}}{\chi_0}\$. When the pressure is changed to 1.50 atm, its volume is \$\frac{250.mt}{\chi_0}\$. What was the original volume?

Skills Worksheet Boyle S Law Key

KJ Lindholm-Leary

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