## 1.2 Mole Conversions Solutions

PRACTICE EXAMPLES. To complete the practice examples (1) write the type of question you have, (2) write the equation with units, (3) show the cancellation of units and (4) calculate the solution.

1. Determine how many atoms are present in 2.5 moles of Silicon? (15.05 x 1023 atoms)

MOLES 
$$\rightarrow$$
 ATOMS 
$$\frac{2.5\_moles}{1} \times \frac{6.022 \times 10^{23} atoms}{1\_mole} = 15.05 \times 10^{23} atoms$$

2. How many moles of Silicon is 3.01 x 10<sup>23</sup> atoms? (.5 moles)

ATOMS 
$$\rightarrow$$
 MOLES 
$$\frac{3.011 \times 10^{23} atoms}{1} \times \frac{1 \_ mole}{6.022 \times 10^{23} atom} = .5 \text{ moles}$$

3. What is the mass of 4 moles of Iron? (224 grams)

MOLES 
$$\rightarrow$$
 GRAMS 
$$\frac{4\_moles}{1} \times \frac{55.9\_grams}{mole} = \frac{224 \text{ grams of C}}{1}$$

4. How many atoms are in 48.6 grams of Magnesium? (12.04 x 1023 atoms)

GRAMS 
$$\rightarrow$$
 ATOMS 
$$\frac{48.6\_grams}{1} \times \frac{1\_mole}{24.3\_grams} \times \frac{6.022 \times 10^{23} atoms}{1\_mole} \approx \frac{12.04 \times 10^{23} atoms}{12.04 \times 10^{23} atoms}$$

5. How many grams are in 18.06 x 1023 atoms of Copper? (190.5 grams)

ATOMS 
$$\rightarrow$$
 GRAMS 
$$\frac{18.06 \times 10^{23} atoms}{1} \times \frac{1\_mole}{6.022 \times 10^{23} atoms} \times \frac{63.5\_grams}{1\_mole} \approx 190.5 \text{ grams}$$

How many moles of Nickel is 176 grams? (3 moles)

GRAMS 
$$\rightarrow$$
 MOLES 
$$\frac{176\_grams\_of\_C}{1} \times \frac{mole}{58.7\_grams} \approx 3 \text{ moles}$$

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**M** Lipman

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