

D - Molar Mass

79. Consider the following:

a. If the formula of a compound is known (e.g., C_4H_4OCl), and you find the masses of the component atoms from the Periodic Table ($C = 12.011g$, $H = 1.0079$, $O = 15.9994$, $Cl = 35.453g$), HOW would you find the molar mass of the compound?

b. Calculate the molar mass of one mole of C_4H_4OCl .

c. Calculate the mass of 5 moles of C_4H_4OCl .

E - Percent Composition of Compounds

80. Provide the word equation describing how you solve the mass percent for a given element in a compound.

90. Penicillin is an important antibiotic that has the chemical formula $C_{16}H_{18}N_2SO_4$. Calculate the mass percent of EACH element in this compound.

F - Empirical Formulas

100. Consider the following:

a. Taking in to account the formula for glucose, $C_6H_{12}O_6$, what information can be gathered about this compound?

b. What ratio of $C:H:O$ is present when considering the formula above?

c. Now take in to account the following formulas: $C_4H_8O_4$, $C_2H_4O_2$, and $C_3H_6O_3$. What do these formulas have in common with glucose?

113. Using the word "ratio" in the descriptions, explain what the difference is between empirical and

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**Donald J. Wink, Sharon Fetzer-
Gislason, Sheila McNicholas**



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