

D - Molar Mass

79. Consider the following:

a. If the formula of a compound is known (e.g., C_4H_4OCl), and you find the masses of the component atoms from the Periodic Table ($C = 12.011g$, $H = 1.0079$, $O = 15.9994$, $Cl = 35.453g$), HOW would you find the molar mass of the compound?

b. Calculate the molar mass of one mole of C_4H_4OCl .

c. Calculate the mass of 5 moles of C_4H_4OCl .

E - Percent Composition of Compounds

80. Provide the word equation describing how you solve the mass percent for a given element in a compound.

90. Penicillin is an important antibiotic that has the chemical formula $C_{16}H_{18}N_2SO_4$. Calculate the mass percent of EACH element in this compound.

F - Empirical Formulas

100. Consider the following:

a. Taking in to account the formula for glucose, $C_6H_{12}O_6$, what information can be gathered about this compound?

b. What ratio of $C:H:O$ is present when considering the formula above?

c. Now take in to account the following formulas: $C_4H_8O_4$, $C_2H_4O_2$, and $C_3H_6O_3$. What do these formulas have in common with glucose?

113. Using the word "ratio" in the descriptions, explain what the difference is between empirical and

Skill Practice 3answer Key Moles And Formulas

Terry C. Jones



Skill Practice 3answer Key Moles And Formulas:

Decoding **Skill Practice 3answer Key Moles And Formulas**: Revealing the Captivating Potential of Verbal Expression

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