

Modeled after AP 2013 #21:

1. $2\text{C}_2\text{H}_6(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of $\text{CO}_2(\text{g})$ is produced?
 A. 0.50 mol
 B. 1.50 mol
 C. **4.50 mol**
 D. 9.00 mol
2. $\text{C}_2\text{H}_6(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of $\text{CO}_2(\text{g})$ is produced?
 A. 0.50 mol
 B. 1.00 mol
 C. **3.50 mol**
 D. 7.00 mol
3. $2\text{C}_2\text{H}_6(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of products made when 3.50 mole of $\text{O}_2(\text{g})$ is consumed?
 A. 2.00 mol
 B. 3.00 mol
 C. **5.00 mol**
 D. 10.0 mol
4. $\text{C}_2\text{H}_6(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of products made when 3.50 mole of $\text{O}_2(\text{g})$ is consumed?
 A. 1.00 mol
 B. **2.00 mol**
 C. 4.00 mol
 D. 8.00 mol

Modeled after AP 2014g #48:

3. $4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
A 7.0 mol sample of $\text{NH}_3(\text{g})$ and a 10. mol sample of $\text{O}_2(\text{g})$ are placed in a previously evacuated container and react according to the equation above. How many moles of $\text{NO}(\text{g})$ are produced when one of the reactants has been totally consumed?
 A. 2.5 mol
 B. **7.0 mol**
 C. 8.0 mol
 D. 28 mol
6. $\text{C}_2\text{H}_6(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{CH}_4(\text{g})$
Samples containing 4.0 mol $\text{C}_2\text{H}_6(\text{g})$ and 5.0 mol $\text{H}_2(\text{g})$ are placed in a previously evacuated container and allowed to react according to the equation above. When one of the reactants is completely consumed, how much product is formed?

Zumdahl Ap Chemistry Review Questions Answers

Yijin Wang



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Table of Contents Zumdahl Ap Chemistry Review Questions Answers

1. Understanding the eBook Zumdahl Ap Chemistry Review Questions Answers
 - The Rise of Digital Reading Zumdahl Ap Chemistry Review Questions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Zumdahl Ap Chemistry Review Questions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Zumdahl Ap Chemistry Review Questions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Zumdahl Ap Chemistry Review Questions Answers
 - Personalized Recommendations
 - Zumdahl Ap Chemistry Review Questions Answers User Reviews and Ratings
 - Zumdahl Ap Chemistry Review Questions Answers and Bestseller Lists
5. Accessing Zumdahl Ap Chemistry Review Questions Answers Free and Paid eBooks
 - Zumdahl Ap Chemistry Review Questions Answers Public Domain eBooks
 - Zumdahl Ap Chemistry Review Questions Answers eBook Subscription Services
 - Zumdahl Ap Chemistry Review Questions Answers Budget-Friendly Options
6. Navigating Zumdahl Ap Chemistry Review Questions Answers eBook Formats

- ePub, PDF, MOBI, and More
- Zumdahl Ap Chemistry Review Questions Answers Compatibility with Devices
- Zumdahl Ap Chemistry Review Questions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Zumdahl Ap Chemistry Review Questions Answers
 - Highlighting and Note-Taking Zumdahl Ap Chemistry Review Questions Answers
 - Interactive Elements Zumdahl Ap Chemistry Review Questions Answers
- 8. Staying Engaged with Zumdahl Ap Chemistry Review Questions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Zumdahl Ap Chemistry Review Questions Answers
- 9. Balancing eBooks and Physical Books Zumdahl Ap Chemistry Review Questions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Zumdahl Ap Chemistry Review Questions Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Zumdahl Ap Chemistry Review Questions Answers
 - Setting Reading Goals Zumdahl Ap Chemistry Review Questions Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Zumdahl Ap Chemistry Review Questions Answers
 - Fact-Checking eBook Content of Zumdahl Ap Chemistry Review Questions Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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