

# Topic 8: Acids and bases

## 8.1 Theories of acids and bases

- 8.1.1 A Brønsted-Lowry acid is a proton/ $H^+$  donor and a Brønsted-Lowry base is a proton/ $H^+$  acceptor
- 8.1.2 Amphiprotic species can act as both Brønsted-Lowry acids and bases
- 8.1.3 A pair of species differing by a single proton is called a conjugate acid-base pair
- 8.1.4 Deduction of the Brønsted-Lowry acid and base in a chemical reaction
- 8.1.5 Deduction of the conjugate acid or conjugate base in a chemical reaction

### Brønsted-Lowry theory

- According to the Brønsted-Lowry theory, a substance behaves as an acid when it donates a proton to a base. A substance behaves as a base when it accepts a proton from an acid. Thus:
  - A Brønsted-Lowry acid is a proton ( $H^+$ ) donor
  - A Brønsted-Lowry base is a proton ( $H^+$ ) acceptor
- Acids can be a combination of **hydrogen ions ( $H^+$ )** and an anion. Examples include:  $HCl$ ,  $HNO_3$ ,  $HC_2H_3O_2$
- Bases can be a combination of **hydroxide ions ( $OH^-$ )** and metal cations. Examples include:  $NaOH$ ,  $KOH$ 
  - However, sometimes a hydrogen next to a metal signifies a base such as  $NaH$
- Acid-Base Conjugate Pairs: A conjugate pair is two species which differ by a single proton:
  - The acid will become the conjugate base
  - The base will become the conjugate acid

#### Conjugate-Base Pair Question:

Label the conjugate acid-base pairs in the following reactions:



Conjugate pair 1:  $CH_3COOH/CH_3COO^-$

$CH_3COOH$  is the acid because it donates a  $H^+$ , so  $CH_3COO^-$  is conjugate base

Conjugate pair 2:  $H_2O/H_3O^+$

$H_2O$  is the base because it accepts a  $H^+$ , so  $H_3O^+$  is the conjugate acid

- To find the conjugate acid of a species, add one  $H^+$
- To find the conjugate base of a species, take away one  $H^+$
- Before the Brønsted-Lowry theory acids and bases were distinguished by their taste, acids tasted sour

### Amphiprotic

- Some substances can behave as either acids and bases, depending on what they react with, and can therefore donate or receive protons. Such substances are said to be amphiprotic
  - **Amphiprotic:** A chemical species capable of accepting and donating protons, thus able to act as a Brønsted-Lowry acid and a base
- This means the species must be able to accept or donate a proton to another species
- $H_2O$  is amphiprotic as it is able to donate a proton to form a  $OH^-$  or accept a proton to form  $H_3O^+$  therefore acting as a Brønsted-Lowry acid or base



#### Amphiprotic Species Question:

Write an equation to show hydrogen phosphate acting as and an acid and a base in water

Acting like an Acid:  $HPO_4^{2-} + H_2O(l) \rightleftharpoons PO_4^{3-} + H_3O^+$

Acting like a Base:  $HPO_4^{2-} + H_2O(l) \rightleftharpoons H_2PO_4^- + OH^-$

# The Physical Setting Chem Topic 8

**Robert Friis**



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