

1. What explains the very high melting and boiling point of water

- | | |
|--|--|
| <input type="radio"/> A Strong dipole-dipole bonds between water molecules | <input type="radio"/> B Dispersion forces which are present in all molecules |
| <input type="radio"/> C Strong hydrogen bonds between water molecules | <input type="radio"/> D Asymmetrical shape of the polar bonds. |

2. Does HCl have hydrogen bonding?

- | | |
|----------------------------|-----------------------------|
| <input type="radio"/> A no | <input type="radio"/> B yes |
|----------------------------|-----------------------------|

3. Does CH₄ have hydrogen bonding?

- | | |
|-----------------------------|----------------------------|
| <input type="radio"/> A yes | <input type="radio"/> B no |
|-----------------------------|----------------------------|

4. All molecules have London forces between them, but dipole-dipole and hydrogen bonding are so much stronger that when they are present we can ignore London forces. Which of these has ONLY London forces?

- | | |
|---|--|
| <input type="radio"/> A I ₂ | <input type="radio"/> B OCl ₂ |
| <input type="radio"/> C NH ₃ | <input type="radio"/> D SH ₂ |

5. Which substance would have the weakest intermolecular forces of attraction?

- | | |
|---|--|
| <input type="radio"/> A CH ₄ | <input type="radio"/> B H ₂ O |
| <input type="radio"/> C NaCl | <input type="radio"/> D MgF ₂ |

6. Intermolecular forces for: NH₃



- | | |
|---------------------------------------|--|
| <input type="radio"/> A Dipole dipole | <input type="radio"/> B Dispersion Force |
|---------------------------------------|--|

States Of Matter And Intermolecular Forces Test

**David Shubert, David C Shubert, John
Leyba**



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States Of Matter A. Goel,2006 This book has been written for the students of under graduate and post graduate level of the various universities in India A special feature of the book is that the text has been illustrated with a large number of line diagrams and the data presented in the form of numerous tables for reference and comparison In the preparation of text standard works and review by renowned author have been freely consulted and the reference given chapter wise At the end of the book will be found useful by those who wish to make a more detailed study of the topics discussed Contents Liquid State Solids State Gaseous State

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