

Heat of Formation Worksheet

Use a standard enthalpies of formation table to determine the change in enthalpy for each of these reactions.



$$[(-411.0) + (-241.8)] - [(-426.7) + (-92.3)] = -133.8 \text{ kJ}$$



$$[2(-393.5)] - [2(-110.5)] = -566 \text{ kJ}$$



$$[(-393.5) + 2(-285.8)] - [(-74)] = -891.1 \text{ kJ}$$



$$[2(-285.8) + 2(-296.1)] - [2(-20.1)] = -1123.6 \text{ kJ}$$



$$[2(+33.9)] - [2(+90.4)] = -113 \text{ kJ}$$

Compound	ΔH_f (kJ/mol)	Compound	ΔH_f (kJ/mol)
$\text{CH}_4\text{(g)}$	-74.8	HCl(g)	-92.3
$\text{CO}_2\text{(g)}$	-393.5	$\text{H}_2\text{O(g)}$	-241.8
NaCl(s)	-411.0	$\text{SO}_2\text{(g)}$	-296.1
$\text{H}_2\text{O(l)}$	-285.8	$\text{NH}_4\text{Cl(s)}$	-315.4
$\text{H}_2\text{S(g)}$	-20.1	NO(g)	+90.4
$\text{H}_2\text{SO}_4\text{(l)}$	-811.3	$\text{NO}_2\text{(g)}$	+33.9
$\text{MgSO}_4\text{(s)}$	-1278.2	$\text{SnCl}_4\text{(l)}$	-545.2
MnO(s)	-384.9	SnO(s)	-286.2
$\text{MnO}_2\text{(s)}$	-519.7	$\text{SnO}_2\text{(s)}$	-580.7
NaCl(s)	-411.0	$\text{SO}_3\text{(g)}$	-296.1
NaF(s)	-569.0	$\text{SO}_3\text{(g)}$	-395.2
NaOH(s)	-426.7	ZnO(s)	-348.0
$\text{NH}_3\text{(g)}$	-46.2	ZnS(s)	-202.9

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