

Chemistry 105

Chapter 3 Review Problems

1. How many molecules are present in 180 grams of water? (The atomic weights of H and O are 1 g/mol and 16 g/mol, respectively.)
a) 6.0×10^{24}
b) 10
c) 5.6×10^{-23}
d) 1.8×10^7
e) 180×10^{23}
2. The atomic mass of sulfur is given as 32.07 g/mol. What is the mass of an average atom of sulfur?
a) 5.33×10^{-23} g
b) 32.07 g
c) 32.07×10^{-23} kg
d) 193×10^{-22} g
e) 6.02×10^{23} g
3. What is the mass of C present in 60 g of C_2H_6 ?
a) 12 g
b) 24 g
c) 48 g
d) 60 g
e) 6 g
4. What is the molar mass of C_2H_6O ?
a) 24 g/mol
b) 32 g/mol
c) 7 g/mol
d) 35 g/mol
e) 46 g/mol
5. What is the correct percentage by mass of nitrogen in HCN?
a) 33 %
b) 14%
c) 27%
d) 52%
e) 98%
6. What is the chief danger that is anticipated if CO_2 levels get too high in the atmosphere?
a) people will develop skin cancer from exposure to UV light
b) the atmosphere will develop an unpleasant odor
c) the temperature of the planet may climb by several degrees
d) plant life will be exterminated
e) The CO_2 will remove ozone from the stratosphere

Standardized Test Practice Chapter 3 Chemistry Answers

Jianjun Gao



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Table of Contents Standardized Test Practice Chapter 3 Chemistry Answers

1. Understanding the eBook Standardized Test Practice Chapter 3 Chemistry Answers
 - The Rise of Digital Reading Standardized Test Practice Chapter 3 Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Standardized Test Practice Chapter 3 Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Standardized Test Practice Chapter 3 Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Standardized Test Practice Chapter 3 Chemistry Answers
 - Personalized Recommendations
 - Standardized Test Practice Chapter 3 Chemistry Answers User Reviews and Ratings
 - Standardized Test Practice Chapter 3 Chemistry Answers and Bestseller Lists
5. Accessing Standardized Test Practice Chapter 3 Chemistry Answers Free and Paid eBooks
 - Standardized Test Practice Chapter 3 Chemistry Answers Public Domain eBooks
 - Standardized Test Practice Chapter 3 Chemistry Answers eBook Subscription Services

- Standardized Test Practice Chapter 3 Chemistry Answers Budget-Friendly Options
- 6. Navigating Standardized Test Practice Chapter 3 Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Standardized Test Practice Chapter 3 Chemistry Answers Compatibility with Devices
 - Standardized Test Practice Chapter 3 Chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Standardized Test Practice Chapter 3 Chemistry Answers
 - Highlighting and Note-Taking Standardized Test Practice Chapter 3 Chemistry Answers
 - Interactive Elements Standardized Test Practice Chapter 3 Chemistry Answers
- 8. Staying Engaged with Standardized Test Practice Chapter 3 Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Standardized Test Practice Chapter 3 Chemistry Answers
- 9. Balancing eBooks and Physical Books Standardized Test Practice Chapter 3 Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Standardized Test Practice Chapter 3 Chemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Standardized Test Practice Chapter 3 Chemistry Answers
 - Setting Reading Goals Standardized Test Practice Chapter 3 Chemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Standardized Test Practice Chapter 3 Chemistry Answers
 - Fact-Checking eBook Content of Standardized Test Practice Chapter 3 Chemistry Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
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