

Chemical Equilibrium

Multiple Choice Practice

1. Chemical reactions are in a state of dynamic equilibrium only when

A. the rate of the forward reaction equals that of the backward reaction.
B. the concentration of reactants and products are equal
C. the activation energies of the forward and backward reactions are equal
D. the reaction involves no enthalpy change.

2. Under the conditions used industrially, ethene and steam react as follows:



Which set of conditions would give the best yield of ethanol at equilibrium?

A. high temperature, low pressure
B. high temperature, high pressure
C. low temperature, high pressure
D. low temperature, low pressure.

3. $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$

In the Haber process, illustrated by the equation, the reaction never reaches equilibrium because

A. the ammonia is constantly removed
B. an iron catalyst is used
C. high pressure is used
D. the temperature is held at above 500°C .

4. The equation refers to the preparation of methanol from synthesis gas.



The formation of methanol is favoured by

A. high pressure and low temperature
B. high pressure and high temperature
C. low pressure and low temperature
D. low pressure and high temperature.

5. $\text{ICl}(\text{l}) + \text{Cl}_2(\text{g}) \rightleftharpoons \text{ICl}_3(\text{s})$

$$\Delta H = -46 \text{ kJ mol}^{-1}$$

Which of the following sets of changes will cause the greatest increase in the proportion of solid in the above equilibrium?

	Temperature	Pressure
A	decrease	decrease
B	decrease	increase
C	increase	decrease
D	increase	increase

6. When a reversible chemical reaction is at equilibrium,

A. the concentration of reactants and products remain equal
B. the forward reaction is unable to continue
C. the concentration of reactants and products remain constant
D. the forward and reverse reactions proceed at different rates.

Static Equilibrium Multiple Choice Questions

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