

Answers to self-tests and exercises

Please note that just the final numerical solutions and short-answer solutions are supplied. For full worked solutions, please refer to the Solutions Manual by Alex Hadenic (97804706781712).

CHAPTER 1

Self-tests

- ST.1 7.9987 am^{-1} , $1.28 \times 10^{-17} \text{ m}$ or 1280 nm .
- ST.2 3d set of orbitals, 5 orbitals
- ST.3 d
- ST.4 3p
- ST.5 The s electron already present in Li repels the incoming electron more strongly than the p electron already present in B repels the incoming p electron, because the incoming p electron goes into a new orbital.
- ST.6 $\text{Ni} : [\text{Ar}][\text{d}^8 4s^2]$, $\text{Ni}^{2+} : [\text{Ar}][\text{d}^8]$
- ST.7 Period 4, Group 2, s block
- ST.8 Generally, going down a group the atomic radius increases and the first ionization energy generally decreases.
- ST.9 Group 14.
- ST.10 The electron-gain process for N is unfavourable.
- ST.11 Na^+

Exercises

- E.1 $J(1s^2)(2s2p^2) = J^2(2s^2)(2p^2) = 2^2/4! = 0.25$
- E.2 (a) 14 nucleons
(b) Exactly at Bohr radius, a_0
(c) $3a_0 \sqrt{5}/4$
- E.3 $J(2p, n=1) = J(2p, n=4) = 13.6 \text{ eV} = 0.378 \text{ aV} = 13.2 \text{ aV}$.
- E.4 0.544 eV . The discrepancy is due to the shielding effect.
- E.5 14.0 eV , 4.16 eV

- E.6 102.5 nm , $1.937 \times 10^{-18} \text{ J}$
- E.7 $1.526 \times 10^{-15} \text{ nm}^{-1}$, 656.3 nm
- E.8 $\frac{1}{\lambda} = R \left(\frac{1}{1^2} - \frac{1}{\infty^2} \right) = 1.0974 \times 10^7 \text{ m}^{-1}$
 $\frac{1}{\lambda} = R \left(\frac{1}{1^2} - \frac{1}{4^2} \right) = 1.0288 \times 10^7 \text{ m}^{-1}$
 $\frac{1}{\lambda} = R \left(\frac{1}{1^2} - \frac{1}{9^2} \right) = 9.7547 \times 10^6 \text{ m}^{-1}$
 $\frac{1}{\lambda} = R \left(\frac{1}{1^2} - \frac{1}{25^2} \right) = 8.2305 \times 10^6 \text{ m}^{-1}$
- E.9 For a given value of n , the angular momentum quantum number l can assume all integer values from 0 to $n-1$.
- E.10 m^2 (a.g., $m^2 = 0$ for $n = 1$, $m^2 = 4$ for $n = 2$, etc.)

E.11

n	l	m_l	Orbital designation	Number of orbitals
1	0	+1, 0, -1	2p	3
2	1	+2, -1, -2	3d	5
		+1, -1	4s	1
3	2	+3, -2, -3	4f	7
		+2, -2		

- E.12 When $n = 5$, $l = 0$ (for the 1 orbitals) and $m_l = -3, -2, -1, 0, 1, 2, 3$, which represent the seven orbitals that complete the 5d subshell. The 5d orbitals represent the start of the actinoids, starting with Th and ending with La.
- E.13 Comparing the plots for 1s (Figure 1.10b) and 2s (Figure 1.12) orbitals, the radial distribution function for a 1s orbital has a single maximum, and that for a 2s orbital has two maxima and a minimum (at $r = 2a_0/2$ for hydrogenic 2s orbitals). The presence of the node at $r = 2a_0/2$ for $40(2s)$ requires the presence of the two maxima and the minimum in the 2s radial distribution functions. The absence of a radial node for $40(2p)$ requires that the 2p radial distribution functions has only a single maximum.

Solutions Manual Inorganic Chemistry Huheey

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