

PROBLEM 2-37

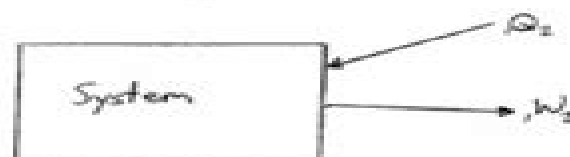
GIVEN: A system of fixed mass at constant pressure and temperature

FIND: Equivalent system form of Eqn. 2.35 (First Law) which is used to define the heat of reaction.

ASSUMPTIONS: expansion/contraction of the system boundary is a reversible process; only work done by system is boundary expansion/contraction

APPROACH: Write the first law for the system. Substitute the appropriate expression for constant pressure work and solve for the specific heat transfer.

Sketch of system:



first law for a system:

$$Q_2 - W_2 = m(u_2 - u_1)$$

for a system at constant pressure the reversible work can be expressed as

$$W_2 = \int_1^2 P dV = P(V_2 - V_1) = mP(v_2 - v_1) \quad v \equiv \text{specific volume}$$

substituting into the first law

$$Q_2 - mP(v_2 - v_1) = m(u_2 - u_1)$$

$$Q_2 = m[(u_2 + Pv_2) - (u_1 + Pv_1)] = m(h_2 - h_1)$$

$$q_2 = Q_2/m = h_2 - h_1$$

if state 1 is the reactants and state 2 is the products then

$$q = h_p - h_r \quad \leftarrow \text{same form as Eqn 2.35}$$

COMMENTS: Note that the first law for a constant pressure system has the same form as for a control volume. This is only true for constant pressure. Also, the internal energies u_1 and u_2 are mixture internal energies. Even assuming ideal gas behavior, $u_1 \neq u_2$ at constant temperature if the composition changes. (see Joule's experiment with gases and internal energy)

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