

Chem 1020 Molarity worksheet, version 2

1. What is the molarity of a solution that contains 5.83 g LiCl in 1.50 L solution?

$$5.83 \text{ g LiCl} \times \frac{1 \text{ mol LiCl}}{42.389 \text{ g LiCl}} = 0.13754 \text{ mol LiCl}$$

$$M = \frac{\text{mol solute}}{\text{L solution}} = \frac{0.13754 \text{ mol LiCl}}{1.50 \text{ L solution}} = 0.0917 \text{ M}$$

2. What is the volume of a 0.300 M CuSO₄ solution that contains 12.4 g CuSO₄?

$$12.4 \text{ g CuSO}_4 \times \frac{1 \text{ mol CuSO}_4}{159.6 \text{ g CuSO}_4} \times \frac{1 \text{ L solution}}{0.300 \text{ mol CuSO}_4} = 0.26 \text{ L solution}$$

3. What mass of solid KBr would be needed to prepare 750.0 mL of 1.75 M KBr solution? How specifically would this solution be prepared?

$$750.0 \text{ mL KBr soln} \times \frac{1 \text{ L soln}}{1000 \text{ mL soln}} \times \frac{1.75 \text{ mol KBr}}{1 \text{ L soln}} \times \frac{119.01 \text{ g KBr}}{1 \text{ mol soln}} = 156 \text{ g KBr}$$

Prepare this solution by measuring 156 g KBr into a 750 mL volumetric flask, and adding enough water to make 750 mL solution.

4. What is the % by mass of 28.3 g of solution which contains 3.56 g NaCl?

$$\% (\text{mass}) = \frac{\text{mass solute}}{\text{mass solution}} \times 100\% = \frac{3.56 \text{ g NaCl}}{28.3 \text{ g solution}} \times 100\% = 12.6\%$$

5. What volume of methanol is required to prepare 2.0 L of a 15.0% by volume solution?

$$\% (\text{vol}) = \frac{\text{vol solute}}{\text{vol solution}} \times 100\%$$

$$\text{vol solute} = \left(\frac{\% (\text{vol})}{100\%} \right) \times \text{vol solution} = \left(\frac{15.0\%}{100\%} \right) \times 2.0 \text{ L} = 0.30 \text{ L}$$

6. What volume of 12.0 M HCl is required to make 500.0 mL of 0.400 M HCl? How specifically would this solution be prepared?

$$M_1 V_1 = M_2 V_2$$

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(0.400 \text{ M})(500.0 \text{ mL})}{(12.0 \text{ M})} = 16.7 \text{ mL}$$

Prepare this solution by measuring 16.7 mL of 12.0 M HCl into a 500 mL volumetric flask, and adding water to make 500 mL of solution.

7. If 17.0 mL of 16.0 M H₂SO₄ solution is diluted to 100.0 mL, what is the diluted concentration?

$$M_1 V_1 = M_2 V_2$$

$$M_2 = \frac{M_1 V_1}{V_2} = \frac{(16.0 \text{ M})(17.0 \text{ mL})}{(100.0 \text{ mL})} = 2.72 \text{ M}$$

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The Princeton Review



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General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition. Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works. This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value. *Mathematics for the Clinical Laboratory E-Book* Lorraine J. Doucette, 2020-06-18 Selected for Doody's Core Titles 2024 with Essential Purchase designation in Laboratory Technology Master the skills you'll need to perform accurate clinical laboratory calculations. *Mathematics for the Clinical Laboratory* 4th Edition demonstrates the calculations used in the analysis of test specimens. It begins by explaining basic mathematical principles and then covers the types of calculations needed in specific areas of the clinical lab including urinalysis, hematology and microbiology. Finally it focuses on the statistical calculations used in quality assurance and quality control. Step by step examples reinforce your understanding and calculation templates and practice problems ensure that you make correct calculations every time. Step by step examples explain basic mathematical principles and show you exactly how to perform each type of calculation. Sample problems with answers can also be used as templates for solving laboratory calculations. Practice problems at the end of each chapter provide a self assessment tool helping you determine what you need to review. Summaries of important formulas are included at the end of the text's major sections. Coverage of statistical calculations includes standard deviation as well as calculations associated with quality assurance and quality control. Quick tips and notes make it easier to understand and remember pertinent information. Learning objectives at the beginning of each chapter provide measurable outcomes to achieve by completing the chapter material. Full color design includes 100 illustrations. Useful appendix of Greek symbols provides a quick reference to turn to when studying. Glossary at the back of the textbook includes definitions of important mathematical terms. New Updated content and calculations reflect the latest procedures used in today's laboratories. **High School Chemistry Unlocked** The Princeton Review, 2016-10-18 UNLOCK THE SECRETS OF CHEMISTRY with THE PRINCETON REVIEW High School Chemistry Unlocked focuses on giving you a wide range of key lessons to help increase your understanding of chemistry. With this book you'll move from foundational concepts to complicated real world applications building confidence as your skills improve. End of chapter drills will help test your comprehension of each facet of chemistry from atoms to alpha radiation. Don't feel locked out. Everything You Need to Know About Chemistry. Complex concepts explained in straightforward ways. Walk throughs of sample problems for all topics. Clear goals and self assessments to help you pinpoint areas for further review. Guided examples of how to solve problems for common subjects. Practice Your Way to Excellence 165 hands on practice questions seeded throughout the chapters and online. Complete answer explanations to boost understanding. Bonus online questions similar to those you'll find on the AP Chemistry Exam and the SAT Chemistry Subject Test. High School Chemistry Unlocked covers Building blocks of matter.

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Table of Contents Solutions Molarity And Dilution Practice Answer Key

1. Understanding the eBook Solutions Molarity And Dilution Practice Answer Key
 - The Rise of Digital Reading Solutions Molarity And Dilution Practice Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Solutions Molarity And Dilution Practice Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solutions Molarity And Dilution Practice Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solutions Molarity And Dilution Practice Answer Key
 - Personalized Recommendations
 - Solutions Molarity And Dilution Practice Answer Key User Reviews and Ratings
 - Solutions Molarity And Dilution Practice Answer Key and Bestseller Lists
5. Accessing Solutions Molarity And Dilution Practice Answer Key Free and Paid eBooks
 - Solutions Molarity And Dilution Practice Answer Key Public Domain eBooks
 - Solutions Molarity And Dilution Practice Answer Key eBook Subscription Services
 - Solutions Molarity And Dilution Practice Answer Key Budget-Friendly Options
6. Navigating Solutions Molarity And Dilution Practice Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Solutions Molarity And Dilution Practice Answer Key Compatibility with Devices
 - Solutions Molarity And Dilution Practice Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solutions Molarity And Dilution Practice Answer Key
 - Highlighting and Note-Taking Solutions Molarity And Dilution Practice Answer Key
 - Interactive Elements Solutions Molarity And Dilution Practice Answer Key

8. Staying Engaged with Solutions Molarity And Dilution Practice Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solutions Molarity And Dilution Practice Answer Key
9. Balancing eBooks and Physical Books Solutions Molarity And Dilution Practice Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solutions Molarity And Dilution Practice Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solutions Molarity And Dilution Practice Answer Key
 - Setting Reading Goals Solutions Molarity And Dilution Practice Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solutions Molarity And Dilution Practice Answer Key
 - Fact-Checking eBook Content of Solutions Molarity And Dilution Practice Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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