

**PAPER-2 [Code – 5]**  
**IITJEE 2011**  
**PART - I: CHEMISTRY**

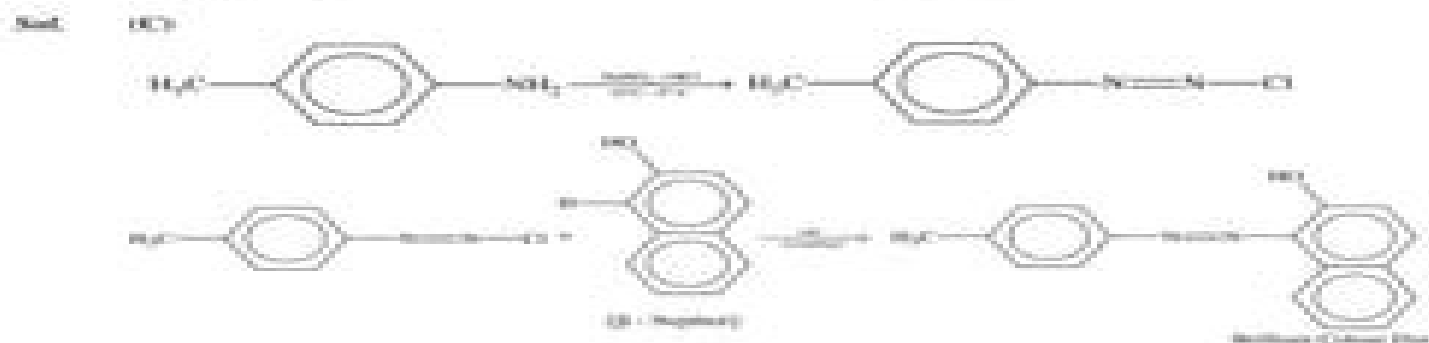
**SECTION – I (Total Marks : 24)**  
**(Single Correct Answer Type)**

This Section contains 8 multiple choice questions. Each question has four choices (A), (B), (C) and (D) out of which **ONLY ONE** is correct.

1. The freezing point (in °C) of a solution containing 0.1 g of  $K_3[Fe(CN)_6]$  (Mol. Wt. 329) in 100 g of water ( $K_f = 1.86 \text{ K kg mol}^{-1}$ ) is  
 (A)  $-2.3 \times 10^{-1}$  (B)  $-5.7 \times 10^{-1}$   
 (C)  $-5.7 \times 10^{-2}$  (D)  $-1.2 \times 10^{-1}$

**Sol.** (A)  
 $K_3[Fe(CN)_6] \rightarrow 3K^+ + [Fe(CN)_6]^{3-}$   
 $i = 4$   
 $\Delta T_f = K_f \times i \times \frac{m}{M} \times \frac{1000}{W} = 1.86 \times 4 \times \frac{0.1}{329} \times \frac{1000}{100} = 2.3 \times 10^{-1}$   
 $T_f = -2.3 \times 10^{-1}$

2. Amongst the compounds given, the one that would form a brilliant colored dye on treatment with  $NaNO_2$  in dil. HCl followed by addition to an alkaline solution of  $\beta$ -naphthol is



# Solved Paper 28 10 2011 Solved Paper

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