



# **ZIMBABWE SCHOOL EXAMINATIONS COUNCIL**

General Certificate of Education Advanced Level

## **BUSINESS STUDIES**

PAPER 1

**6025/1**

**SPECIMEN PAPER**

3 hours

Additional materials:  
Answer paper

**TIME** 3 hours

### **INSTRUCTIONS TO CANDIDATES**

Write your name, centre number and candidate number in the spaces provided on the answer paper/answer booklet.

Answer **all** questions.

Write your answers on the separate answer paper provided.

If you use more than one sheet of paper, fasten the sheets together.

### **INFORMATION FOR CANDIDATES**

The number of marks is given in brackets [ ] at the end of each question or part question.

You are advised to spend not more than 45 minutes on Section A.

The businesses described in this paper are entirely fictitious.

You are reminded of the need for good English and clear presentation in your answers.

Calculators may be used.

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# Zimsec November 2013 Business Studies Paper 1

**Terry C. Jones**



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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h<sub>2</sub>o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^-$   $14\text{H}^+$   $\text{Cr}_2\text{O}_7^{2-} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + \text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + 2\text{e}^-$  equalize the electrons  $6\text{e}^-$   $14\text{H}^+$   $\text{Cr}_2\text{O}_7^{2-} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + 6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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