

Modeled after AP 2013 #21:

1. $2\text{C}_2\text{H}_6(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of $\text{CO}_2(\text{g})$ is produced?
 A. 0.50 mol
 B. 1.50 mol
 C. **4.50 mol**
 D. 9.00 mol
2. $\text{C}_2\text{H}_6(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of $\text{CO}_2(\text{g})$ is produced?
 A. 0.50 mol
 B. 1.00 mol
 C. **3.50 mol**
 D. 7.00 mol
3. $2\text{C}_2\text{H}_6(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of products made when 3.50 mole of $\text{O}_2(\text{g})$ is consumed?
 A. 2.00 mol
 B. 3.00 mol
 C. **5.00 mol**
 D. 10.0 mol
4. $\text{C}_2\text{H}_6(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of products made when 3.50 mole of $\text{O}_2(\text{g})$ is consumed?
 A. 1.00 mol
 B. **2.00 mol**
 C. 4.00 mol
 D. 8.00 mol

Modeled after AP 2014g #48:

3. $4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
A 7.0 mol sample of $\text{NH}_3(\text{g})$ and a 10. mol sample of $\text{O}_2(\text{g})$ are placed in a previously evacuated container and react according to the equation above. How many moles of $\text{NO}(\text{g})$ are produced when one of the reactants has been totally consumed?
 A. 2.5 mol
 B. **7.0 mol**
 C. 8.0 mol
 D. 28 mol
6. $\text{C}_2\text{H}_6(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{CH}_4(\text{g})$
Samples containing 4.0 mol $\text{C}_2\text{H}_6(\text{g})$ and 5.0 mol $\text{H}_2(\text{g})$ are placed in a previously evacuated container and allowed to react according to the equation above. When one of the reactants is completely consumed, how much product is formed?

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