

36. Calculate the mass of Al_2S_3 produced when 100 mL of 0.50 M AlCl_3 is added to 100 mL of 0.50 M Na_2S .
37. You are given the equation: $\text{AgBr} + 2\text{S}_2\text{O}_3^{2-} \rightarrow \text{Ag}(\text{S}_2\text{O}_3)_2^{2-} + \text{Br}^-$. What mass of AgBr can be dissolved by 750. mL of 0.300 M $\text{Na}_2\text{S}_2\text{O}_3$?
38. Given the following chemical equation, determine the theoretical yield of B_2H_6 if exactly 100.0 g of LiAlH_4 was allowed to react with 225 g of BF_3 .
- $$3\text{LiAlH}_4 + 4\text{BF}_3 \rightarrow 3\text{LiF} + 3\text{AlF}_3 + 2\text{B}_2\text{H}_6$$

Section 4.8

39. What volume of 0.1379 M HCl is required to neutralize 10.0 mL of 0.2789 M NaOH solution?
40. How many mL of 1.50 M NaOH is required to neutralize 275 mL of 0.5 M H_2SO_4 ?
41. What is the molarity of a solution of HCl if it requires 29.31 mL of a 0.0923 M NaOH solution to reach a phenolphthalein endpoint for the titration of a 10.0 mL aliquot of the HCl solution?
42. Given the following two equations, determine the number of grams of manganese dioxide required to prepare enough chlorine gas to produce 25.0 g of potassium hypochlorite.
- $$\begin{aligned}\text{MnO}_2 + 4\text{HCl} &\rightarrow \text{MnCl}_2 + \text{Cl}_2 + 2\text{H}_2\text{O} \\ \text{Cl}_2 + 2\text{KOH} &\rightarrow \text{KCl} + \text{KClO} + \text{H}_2\text{O}\end{aligned}$$
43. A titration is done using 0.1302 M NaOH to determine the molar mass of an acid. The acid contains one acidic hydrogen per molecule. If 1.863 g of the acid require 70.11 mL of the NaOH solution, what is the molar mass of the acid?
44. A 2.000-g sample of silver alloy was dissolved in nitric acid and then precipitated as AgBr . After drying, the sample of silver bromide weighed 2.000 g. Calculate the percentage of silver in the alloy.
45. Complete and balance each acid-base equation (assume complete neutralization):
- | | |
|---|---|
| a. $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow$ | c. $\text{H}_2\text{SO}_3 + \text{NaOH} \rightarrow$ |
| b. $\text{H}_3\text{PO}_4 + \text{Mg}(\text{OH})_2 \rightarrow$ | d. $\text{HC}_2\text{H}_3\text{O}_2 + \text{Ba}(\text{OH})_2 \rightarrow$ |
46. What volume of 0.2 M NaOH is required to neutralize 50 mL of 0.1 M H_2SO_4 ?
47. How many mL of 2.3 M HNO_3 is needed to neutralize 0.92 L of 0.5 M KOH ?
48. The neutralization of a 25.00-mL sample of an unknown base requires 18.34 mL of 0.100 M HCl . Assuming that the acid-base stoichiometry is 1:1, what is the concentration of the unknown base?
49. Pennies made after 1982 contain about 97% zinc by mass. A student wants to prove this by filing the copper outside of a penny until he sees zinc and then putting the penny in a 1.00 M HCl solution. The zinc will be oxidized, and the H^+ (from HCl) will be reduced:
- $$\text{Zn} + 2\text{H}^+ \rightarrow \text{Zn}^{2+} + \text{H}_2$$
- If the entire penny has a mass of 2.80 grams, how many mL of 1.00 M HCl are required to just react with all the zinc? (You would, in reality, add much more to completely surround the penny.)
50. Assuming that the stoichiometry is 1:1, what is the concentration of an unknown acid if a 20.0-mL sample of it is neutralized with precisely 33.4 mL of 0.250 M base?

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Susan McMurry, John McMurry



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