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## pH and pOH

The pH of a solution indicates how acidic or basic that solution is.

pH range 0 to 7 acidic

7 neutral

7-14 basic

Since  $[H^+][OH^-] = 10^{-14}$  at 25 °C, if  $[H^+]$  is known, the  $[OH^-]$  can be calculated and vice versa.

$$pH = -\log [H^+]$$

$$pOH = -\log [OH^-]$$

$$\text{So if } [H^+] = 10^{-9} \text{ M, } pH = 9.$$

$$\text{So if } [OH^-] = 10^{-5} \text{ M, } pOH = 5.$$

Together,  $pH + pOH = 14$ .

Complete the following chart.

	$[H^+]$	pH	$[OH^-]$	pOH	Acidic or Basic
1.	$10^{-1} \text{ M}$	1	$10^{-13} \text{ M}$	13	Acidic
2.		2			
3.			$10^{-6} \text{ M}$		
4.	$10^{-8} \text{ M}$				
5.				11	
6.		12			
7.			$10^{-7} \text{ M}$		
8.	$10^{-11} \text{ M}$				
9.				13	
10.		8			

# South Pasadena Ph And Poh Answers

**Rachel S Tattersall**



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