

Solution

Section A

- (d) Molality (m)

Explanation: Molality is defined as no. of moles of solute present per kg of solvent. It is denoted as " m ". Mathematically, Molality (m) = [number of moles of the solute] / kg of solvent
Since this mode of expressing the strength of a solution involves (weight/weight) relationship of solute and solvent, the molality of the solution is not affected by variation in temperatures of the solution.
- (a) behaves like particles

Explanation: The emission of free electrons from a metal surface when light is shone on it is called the photoemission or the photoelectric effect. This effect led to the conclusion that light is made up of packets or quantum of energy. Einstein already associated the light quantum with momentum. This strongly supported the particle nature of light and these particles were named photons. Thus, the wave-particle duality of light came into the picture. Einstein won the Nobel Prize for Physics not for his work on relativity, but for explaining the photoelectric effect.
- (b) $\Delta H = -7.151 \text{ kJ mol}^{-1}$

Explanation: Water (10.0°C) $\rightarrow$ ice (-10.0°C) $\Delta H = ?$
The enthalpy change for the conversion of 1 mole liquid water at 10.0°C into 1 mole liquid water 0°C, water (10.0°C) $\rightleftharpoons$ water(0°C).
 $\Delta H_1 = C_p \times H_2O(l) \times \Delta T = -75.3 \text{ J mol}^{-1} \text{ K}^{-1} = -753 \text{ J mol}^{-1}$
Enthalpy of fusion, water (0°C) $\rightleftharpoons$ ice(0°C), $\Delta H_2 = \Delta_{fus}H$
 $\Delta H_2 = \Delta H_{fusion} = -\Delta H_{freezing} = -6.06 \text{ kJ mol}^{-1}$
Enthalpy change for conversion of 1 mole of ice at 0°C to 1 mole of ice at 10°C, Ice (0°C) $\rightleftharpoons$ ice(10.0°C);
 $\Delta H_3 = C_p[H_2O(s)]\Delta T = 36.8 \text{ J mol}^{-1} \text{ K}^{-1} \times 10\text{K} = 368 \text{ J mol}^{-1}$
 $\Delta H = \Delta H_1 + \Delta H_2 + \Delta H_3 = -(0.753 + 6.03 + 0.368)\text{kJ mol}^{-1} = -7.151 \text{ kJ mol}^{-1}$
- (d) $1.988 \times 10^{-18} \text{ J}$

Explanation: We know Planck's equation is
 $E = h\nu$
where E is energy, h is Planck's constant and ν is frequency.
 $E = 6.626 \times 10^{-34} \times 3 \times 10^{15} = 1.988 \times 10^{-18} \text{ J}$
- (c) 20.15

Explanation: $\Delta G^\circ = -115 \times 10^3 \text{ J}$
 $T = 298 \text{ K}$, $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$
 $-\Delta G^\circ = 2.303 RT \log_{10} K_p$
 $-(-115 \times 10^3) = 2.303 \times 8.314 \times 298 \log_{10} K_p$
 $\log_{10} K_p = \frac{115000}{2.303 \times 8.314 \times 298} = 20.15$
- (d) Paschen series

Explanation: The Lyman series lies in the ultraviolet, whereas the Paschen, Brackett, and Pfund series lies in the infrared.
- (b) reduction

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1. Understanding the eBook Sp 2 Cuaderno Answer Key
 - The Rise of Digital Reading Sp 2 Cuaderno Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Sp 2 Cuaderno Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Sp 2 Cuaderno Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Sp 2 Cuaderno Answer Key
 - Personalized Recommendations
 - Sp 2 Cuaderno Answer Key User Reviews and Ratings

- Sp 2 Cuaderno Answer Key and Bestseller Lists
- 5. Accessing Sp 2 Cuaderno Answer Key Free and Paid eBooks
 - Sp 2 Cuaderno Answer Key Public Domain eBooks
 - Sp 2 Cuaderno Answer Key eBook Subscription Services
 - Sp 2 Cuaderno Answer Key Budget-Friendly Options
- 6. Navigating Sp 2 Cuaderno Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Sp 2 Cuaderno Answer Key Compatibility with Devices
 - Sp 2 Cuaderno Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Sp 2 Cuaderno Answer Key
 - Highlighting and Note-Taking Sp 2 Cuaderno Answer Key
 - Interactive Elements Sp 2 Cuaderno Answer Key
- 8. Staying Engaged with Sp 2 Cuaderno Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Sp 2 Cuaderno Answer Key
- 9. Balancing eBooks and Physical Books Sp 2 Cuaderno Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Sp 2 Cuaderno Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Sp 2 Cuaderno Answer Key
 - Setting Reading Goals Sp 2 Cuaderno Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Sp 2 Cuaderno Answer Key
 - Fact-Checking eBook Content of Sp 2 Cuaderno Answer Key
 - Distinguishing Credible Sources

-
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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