

# N 875

Seat No. 

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2023 III 06 1100 -N 875- ENGLISH (17) (THIRD LANGUAGE) (E)

(REVISED COURSE)

Time : 3 Hours

(Pages 19)

Max. Marks : 80

## (ACTIVITY SHEET)

SECTION I : Language Study – 10 marks

**Q.1.(A) Do as directed (Attempt any four) : (8 marks)**

- (1) Complete the following words by using correct letters : (2)
  - (i) Pe\_ ple
  - (ii) Mat\_h
  - (iii) Hou\_e
  - (iv) Ad\_It
- (2) Put the following words in alphabetical order : (2)
  - (i) survive, leader, discuss, present
  - (ii) light, loyal, luck, large
- (3) Punctuate the following sentences : (2)
  - (i) well i wish you luck. said sitaram
  - (ii) whats your name little one asked the teacher
- (4) Make four words (minimum of three letters each) using the letters in the word : (2)

"Benchmark"

P.T.O.

# Ssc English Paper

**Mocktime Publication**



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web mar 25 2020 this activity supports the following unit topics and learning objectives unit 7 equilibrium topic 7 1 introduction to equilibrium tra 6 a explain the relationship between the occurrence of a reversible chemical or physical process and the establishment of equilibrium to experimental observations

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web write down the equilibrium equation you are investigating using the information provided by the introduction within the

virtual lab activity equilibrium reaction  $\text{CO} + \text{H}_2\text{O} \rightleftharpoons \text{CO}_2 + \text{H}_2$  select each control test tube to compare the concentrations of

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web the study of reaction rates will lead you into the study of chemical equilibrium in this semester you will also study electrochemistry from batteries to electroplating and you will learn about acids bases and salts unit 15 is a brief study of the atomic nucleus which your teacher may introduce at any time during the year

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web equilibrium moles  $\text{HSCN}$   $2.00 \times 10^{-5} \text{ mol}$   $3.00 \times 10^{-6} \text{ mol}$   $1.70 \times 10^{-5} \text{ mol}$   $\text{HSCN}$  knowing the number of moles of  $\text{Fe}^{3+}$  and  $\text{HSCN}$  present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species  $\text{mol Fe}^{3+}$   $1.70 \times 10^{-5} \text{ mol}$   $\text{Fe}^{3+}$   $8.50 \times 10^{-4} \text{ M}$

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web jan 29 2023 7 1 introduction to equilibrium let s start this unit out with a broaaaaad overview of what equilibrium is all about in chemistry we define equilibrium as the point at which the rate of the forward reaction is equal to the rate of the reverse reaction we learned about rates in unit five as how quickly a reaction proceeds

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web equilibrium questions a covalent bond has a standard gibbs free energy of  $400 \text{ kJ mol}^{-1}$  while the hydrogen bond and van der waals forces have substantially lower values respectively  $25$  and  $5 \text{ kJ mol}^{-1}$  given such  $\Delta G$  values which of the following is equivalent to the ratio of the  $K_{eq}$  value of a covalent bond to that of a hydrogen bond

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web nov 13 2022 define the equilibrium state of a chemical reaction system what is its practical significance state the meaning and significance of the following terms reversible reaction quantitative reaction kinetically inhibited reaction explain the meaning of the statement equilibrium is macroscopically static but microscopically dynamic very

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related equipment and processes properly