

HW#5: Solid State Physics

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PHYS-3120

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1. Electrons in Copper

What is the de Broglie wavelength of electrons in copper?

The Fermi-energy of electrons in copper is $E_F = 7.0354$ eV and the de Broglie wavelength is given by the equation:

$$\lambda = \frac{h}{\sqrt{2 m_e E_F}}$$

```
h = Quantity[4.13567 * 10^-15, "Electronvolts" "Seconds"];
me = Quantity[0.511 / c^2, "Megaelectronvolts"];
ef = Quantity[7.0354, "Electronvolts"];
c = Quantity[3 * 10^8, "Meters" "Seconds"^-1];

wl = h / Sqrt[2 * me * ef]

4.62698 * 10^-10 m
```

The de Broglie wavelength of electrons in copper is 0.463 nm.

Compare the de Broglie wavelength to the atomic separation in copper.

The density of copper is 8.94 g cm^{-3} and its mass is $63.546 \text{ g mol}^{-1}$.

We can find the atomic separation of copper by:

$$d = \left(\frac{63.546 \text{ g}}{\text{mol}} \times \frac{\text{cm}^3}{8.94 \text{ g}} \times \frac{1 \text{ mol}}{6.022 \cdot 10^{23}} \right)^{1/3}$$

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