

# STAAR CHEMISTRY REFERENCE MATERIALS



## ATOMIC STRUCTURE

$$\text{Speed of light} = (\text{frequency})(\text{wavelength})$$

$$c = f\lambda$$

$$\text{Energy} = (\text{Planck's constant})(\text{frequency})$$

$$E_{\text{photon}} = hf$$

$$\text{Energy} = \frac{(\text{Planck's constant})(\text{speed of light})}{(\text{wavelength})}$$

$$E_{\text{photon}} = \frac{hc}{\lambda}$$

## BEHAVIOR OF GASES

$$\text{Total pressure of a gas} = \left( \begin{array}{c} \text{sum of the partial pressures} \\ \text{of the component gases} \end{array} \right)$$

$$P_T = P_1 + P_2 + P_3 + \dots$$

$$(\text{Pressure})(\text{volume}) = (\text{moles})(\text{ideal gas constant})(\text{temperature})$$

$$PV = nRT$$

$$\frac{(\text{Initial pressure})(\text{initial volume})}{(\text{Initial moles})(\text{initial temperature})} = \frac{(\text{final pressure})(\text{final volume})}{(\text{final moles})(\text{final temperature})}$$

$$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

$$(\text{Initial pressure})(\text{initial volume}) = (\text{final pressure})(\text{final volume})$$

$$P_1V_1 = P_2V_2$$

$$\frac{(\text{Initial volume})}{(\text{Initial temperature})} = \frac{(\text{final volume})}{(\text{final temperature})}$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{(\text{Initial volume})}{(\text{Initial moles})} = \frac{(\text{final volume})}{(\text{final moles})}$$

$$\frac{V_1}{n_1} = \frac{V_2}{n_2}$$

## SOLUTIONS

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liter of solution}}$$

$$M = \frac{\text{mol}}{\text{L}}$$

$$\text{Ionization constant of water} = \left( \begin{array}{c} \text{hydrogen ion} \\ \text{concentration} \end{array} \right) \left( \begin{array}{c} \text{hydroxide ion} \\ \text{concentration} \end{array} \right)$$

$$K_w = [\text{H}^+][\text{OH}^-]$$

$$\left( \begin{array}{c} \text{Volume of} \\ \text{solution 1} \end{array} \right) \left( \begin{array}{c} \text{molarity of} \\ \text{solution 1} \end{array} \right) = \left( \begin{array}{c} \text{volume of} \\ \text{solution 2} \end{array} \right) \left( \begin{array}{c} \text{molarity of} \\ \text{solution 2} \end{array} \right)$$

$$V_1M_1 = V_2M_2$$

$$\text{pH} = -\log(\text{hydrogen ion concentration})$$

$$\text{pH} = -\log[\text{H}^+]$$

## THERMOCHEMISTRY

$$\text{Heat gained or lost} = (\text{mass}) \left( \begin{array}{c} \text{specific} \\ \text{heat} \end{array} \right) \left( \begin{array}{c} \text{change in} \\ \text{temperature} \end{array} \right)$$

$$Q = mc_p\Delta T$$

$$\text{Enthalpy of reaction} = \left( \begin{array}{c} \text{enthalpy} \\ \text{of products} \end{array} \right) - \left( \begin{array}{c} \text{enthalpy} \\ \text{of reactants} \end{array} \right)$$

$$\Delta H = \Delta H_f^\circ(\text{products}) - \Delta H_f^\circ(\text{reactants})$$

# Staar Chemistry Key

**Xiaolong Qi**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the name bar.

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web objective aims of this study were to identify the association of community public health nursing educators members demographics educational preparation and research competencies and to complete a strengths weaknesses opportunities and threats swot analysis methods a descriptive correlational design was employed

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web methods the swot analysis technique strengths weaknesses opportunities threats was used to assess the current state of advanced practice nursing worldwide with specific focus on regulation education licensure and practice data were collected broadly from peer reviewed governmental and regulatory sources as well as grey literature

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web the results of the swot analysis are typically recorded in a two by two table that shows the four dimensions side by side for comparison as shown in one of the sample swot analyses that follows alternatively the four cells can be presented below each other as is done in the other example