

Gas Law Problems

1. The volume of a gas is 2.32 L at 1.45 atm pressure, at constant temperature, what will be the volume if the pressure is changed to 4.87 atm?

Soln:

2. A gas cylinder is at a pressure of 1500 mmHg at 25° C, it can stand 20.25 atm pressure, at which temperature the cylinder will burst?

Soln:

3. A 438 L steel tank is filled with 0.885 kg O₂, calculate the pressure at 21° C.

Soln:

4. Find density of NH₃ gas at standard temperature and pressure (STP).

Soln:

5. 100 g of each gas N₂, O₂, and CO₂ were mixed in a container at a pressure of 2.6×10^5 Pa. What is the partial pressure of O₂?

Soln:

Boyle's Law

1. At 1.70 atm, a sample of gas takes up 4.25L. If the pressure in the gas is increased to 2.40 atm, what will the new volume be?
2. A sample of Ne gas occupies 0.220L at 0.86 atm. What will be its volume at 29.4kPa?

Charles's Law

1. A balloon takes up 625L at 0°C. If it is heated to 80°C, what will its new volume be?
2. A gas at 40.0°C occupies a volume of 2.32L. If the temperature is raised to 75.0°C, what will the new volume be if the pressure is constant?

Avogadro's Law

Weston Walch Gas Law Problems

J. Weston Walch



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