

# STAAR CHEMISTRY REFERENCE MATERIALS



## ATOMIC STRUCTURE

$$\text{Speed of light} = (\text{frequency})(\text{wavelength}) \quad c = f\lambda$$

$$\text{Energy} = (\text{Planck's constant})(\text{frequency}) \quad E_{\text{photon}} = hf$$

$$\text{Energy} = \frac{(\text{Planck's constant})(\text{speed of light})}{(\text{wavelength})} \quad E_{\text{photon}} = \frac{hc}{\lambda}$$

## BEHAVIOR OF GASES

$$\text{Total pressure of a gas} = \left( \begin{array}{c} \text{sum of the partial pressures} \\ \text{of the component gases} \end{array} \right) \quad P_T = P_1 + P_2 + P_3 + \dots$$

$$(\text{Pressure})(\text{volume}) = (\text{moles})(\text{ideal gas constant})(\text{temperature}) \quad PV = nRT$$

$$\frac{(\text{Initial pressure})(\text{initial volume})}{(\text{Initial moles})(\text{initial temperature})} = \frac{(\text{final pressure})(\text{final volume})}{(\text{final moles})(\text{final temperature})} \quad \frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

$$(\text{Initial pressure})(\text{initial volume}) = (\text{final pressure})(\text{final volume}) \quad P_1V_1 = P_2V_2$$

$$\frac{(\text{Initial volume})}{(\text{Initial temperature})} = \frac{(\text{final volume})}{(\text{final temperature})} \quad \frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{(\text{Initial volume})}{(\text{Initial moles})} = \frac{(\text{final volume})}{(\text{final moles})} \quad \frac{V_1}{n_1} = \frac{V_2}{n_2}$$

## SOLUTIONS

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liter of solution}} \quad M = \frac{\text{mol}}{\text{L}}$$

$$\text{Ionization constant of water} = \left( \begin{array}{c} \text{hydrogen ion} \\ \text{concentration} \end{array} \right) \left( \begin{array}{c} \text{hydroxide ion} \\ \text{concentration} \end{array} \right) \quad K_w = [\text{H}^+][\text{OH}^-]$$

$$\left( \begin{array}{c} \text{Volume of} \\ \text{solution 1} \end{array} \right) \left( \begin{array}{c} \text{molarity of} \\ \text{solution 1} \end{array} \right) = \left( \begin{array}{c} \text{volume of} \\ \text{solution 2} \end{array} \right) \left( \begin{array}{c} \text{molarity of} \\ \text{solution 2} \end{array} \right) \quad V_1M_1 = V_2M_2$$

$$\text{pH} = -\log(\text{hydrogen ion concentration}) \quad \text{pH} = -\log[\text{H}^+]$$

## THERMOCHEMISTRY

$$\text{Heat gained or lost} = (\text{mass}) \left( \begin{array}{c} \text{specific} \\ \text{heat} \end{array} \right) \left( \begin{array}{c} \text{change in} \\ \text{temperature} \end{array} \right) \quad Q = mc_p\Delta T$$

$$\text{Enthalpy of reaction} = \left( \begin{array}{c} \text{enthalpy} \\ \text{of products} \end{array} \right) - \left( \begin{array}{c} \text{enthalpy} \\ \text{of reactants} \end{array} \right) \quad \Delta H = \Delta H_f^\circ(\text{products}) - \Delta H_f^\circ(\text{reactants})$$

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## **Table of Contents Staar Chemistry Formula Chart Scavenger Hunt**

1. Understanding the eBook Staar Chemistry Formula Chart Scavenger Hunt
  - The Rise of Digital Reading Staar Chemistry Formula Chart Scavenger Hunt
  - Advantages of eBooks Over Traditional Books
2. Identifying Staar Chemistry Formula Chart Scavenger Hunt
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Staar Chemistry Formula Chart Scavenger Hunt
  - User-Friendly Interface
4. Exploring eBook Recommendations from Staar Chemistry Formula Chart Scavenger Hunt
  - Personalized Recommendations
  - Staar Chemistry Formula Chart Scavenger Hunt User Reviews and Ratings
  - Staar Chemistry Formula Chart Scavenger Hunt and Bestseller Lists
5. Accessing Staar Chemistry Formula Chart Scavenger Hunt Free and Paid eBooks
  - Staar Chemistry Formula Chart Scavenger Hunt Public Domain eBooks
  - Staar Chemistry Formula Chart Scavenger Hunt eBook Subscription Services
  - Staar Chemistry Formula Chart Scavenger Hunt Budget-Friendly Options
6. Navigating Staar Chemistry Formula Chart Scavenger Hunt eBook Formats

- ePub, PDF, MOBI, and More
  - Staar Chemistry Formula Chart Scavenger Hunt Compatibility with Devices
  - Staar Chemistry Formula Chart Scavenger Hunt Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Staar Chemistry Formula Chart Scavenger Hunt
    - Highlighting and Note-Taking Staar Chemistry Formula Chart Scavenger Hunt
    - Interactive Elements Staar Chemistry Formula Chart Scavenger Hunt
  8. Staying Engaged with Staar Chemistry Formula Chart Scavenger Hunt
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Staar Chemistry Formula Chart Scavenger Hunt
  9. Balancing eBooks and Physical Books Staar Chemistry Formula Chart Scavenger Hunt
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Staar Chemistry Formula Chart Scavenger Hunt
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Staar Chemistry Formula Chart Scavenger Hunt
    - Setting Reading Goals Staar Chemistry Formula Chart Scavenger Hunt
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Staar Chemistry Formula Chart Scavenger Hunt
    - Fact-Checking eBook Content of Staar Chemistry Formula Chart Scavenger Hunt
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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