

1 25.0 cm³ of 0.200 mol/dm³ sulfuric acid neutralises 18.6 cm³ of barium by droxide solution.

H₂SO_{*}(aq) + Ba(OH0₂(aq) → BaSO_{*}(s) + 2H₂O(t)

a Find the concentration of the barium hydroxide solution in mordan?. Give your answer to 3 significant figures.

moles $H_1SO_4 = \text{conc} \times \text{vol} (\text{dm}^2) = 0.200 \times \frac{15.8}{1848} = 0.00500 \text{ mol}$ moles $Ba(OH)_2 = \text{moles of } H_2SO_4 = 0.00500 \text{ mol}$ $\text{conc } Ba(OH)_2 = \frac{\text{moles}}{\text{volume} (\text{dm}^2)} = \frac{0.00500}{184} = 0.269 \text{ mol/dm}^2$

b Find the concentration of the barium hydroxide solution in glidm³. Give your answer to 3 significant figures.

conc Ba(OH) - 171 x 0.269 - 46,0 g/dm²

c List the key steps in the practical you would carry out to find out that 18.6 cm² of the barium hydroxide solution is needed to neutralise 25.0 cm² of 0.200 movidm² sulfuric acid.

use a pipette to place 25 cm² of sulfuric acid in a conical flask add an indicator add barium hydroxide from burette until colour changes go dropwise near end-point record result

2 What volume of 0.0100 mol/dm³ calcium hydroxide neutralises 25.0 cm³ of 0.140 mol/dm³ sulfuric acid? Give your answer to 3 significant figures

H₂SO₄(aq) + Ca(OH)₂(aq) → CaSO₄(s) + 2 H₂O₃()
moles H₂SO₄ = conc × vol (dm²) = 0.140 × 15.8 / 1846 = 0.00350 mol
moles Ca(OH)₂ = moles H₂SO₄ = 0.00350 mol
volume Ca(OH)₂ = moles / 0.00350 = 0.350 dm²

3 What volume of 1.50 mol/dm³ potassium hydroxide reacts with 25.0 cm³ of 1.10 mol/dm³ nitric acid? Give your answer to 3 significant figures

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C Cleary

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