



**WEEK 02 CLASSES:**

# **HYPERMEDIA**



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# Tma4 Cit 735 Test

**VM Jensen**



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web feb 12 2020 tl dr too long didn't read to calculate the amount of heat released in a chemical reaction use the equation  $q = mc \Delta t$  where  $q$  is the heat energy transferred

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two surfaces with different temperatures placed opposite of each other will exchange heat to each other by thermal radiation the heat radiated from the hot surface will be higher see more

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web if it takes an amount  $q$  of heat to cause a temperature change  $\Delta t$  in a given mass of copper it will take 10 8 times that amount of heat to cause

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convection is not the only heat flow from a surface as thermal radiation will radiate energy as electromagnetic waves at the speed of light the percentage of the entire heat flow due to thermal radiation increases with see more

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conduction also known as thermal conduction is the transfer of internal energy within a material the energy is transferred by collision of molecules atoms and electrons inside the material the heat flow will occur in solid liquid and plasma phases and the energy will always flow from hot to cold joseph see more

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web when the reactants have more potential energy than the products have and the extra energy is released to the

surroundings it is called activation energy the height of the hill on a potential energy diagram is a measure of the endothermic the reaction above can be categorized as

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web this is like a one dimensional system whose mechanical energy  $e$  is a constant and whose potential energy with respect to zero energy at zero displacement from the spring  $s$  unstretched length  $x_0$  is  $\frac{1}{2} kx^2$  figure 8 5 2 8 5 2 a a glider between springs on an air track is an example of a horizontal mass spring system

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web 1 if the  $pe$  of the reactants is 430kJ and the  $pe$  of the products is 190kJ what is the  $\Delta h$  2 if the  $pe$  of the reactants is 430kJ and the activation energy is 220 how much total energy will be needed to be added for the reaction to proceed transcribed image text potential energy diagram answer the questions using the graph

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web figure 5 5 1 a potential energy diagram shows the total potential energy of a reacting system as the reaction proceeds a in an endothermic reaction the energy of the products is greater than the energy of the reactants and  $\Delta h$  is positive b in an exothermic reaction the energy of the products is lower than the energy of the

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web answer see below explanation the rate of a chemical reaction  $a + b \rightarrow c + d$  is given in first approximation by  $v = k \frac{[a]^m [b]^n}{[c]^p [d]^q}$  where  $k$  depends by many factors as solvent geometry type of reaction  $E_a$  is the activation energy  $R$  is the universal constant of gases  $T$  is the absolute temperature

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web mar 17 2016 3 what is the activation energy of a reaction and how is this energy related to the activated complex of the reaction  $E_a$  is the minimum amount of energy for a reaction to occur it is the amount of energy required to create an activated complex 4 what happens when a catalyst is used in a reaction a catalyst changes the reaction

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web key points potential energy is energy that has the potential to become another form of energy an object's potential energy depends on its physical properties and position in a system potential energy comes in many forms such as gravitational potential energy due to an object's mass and position in a gravitational field

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web the energy of the reactants is shown on the energy diagram above on the right side of the energy diagram draw a horizontal line segment to indicate the energy of the products draw a vertical line segment to indicate  $\Delta H$  for the reaction and label it with the correct value see attached image

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web may 1 2013 a potential energy diagram shows the change in potential energy of a system as reactants are converted into products figure below shows basic potential energy diagrams for an endothermic a and an exothermic b reaction recall that the enthalpy change  $\Delta H$  is positive for an endothermic reaction and negative for an

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