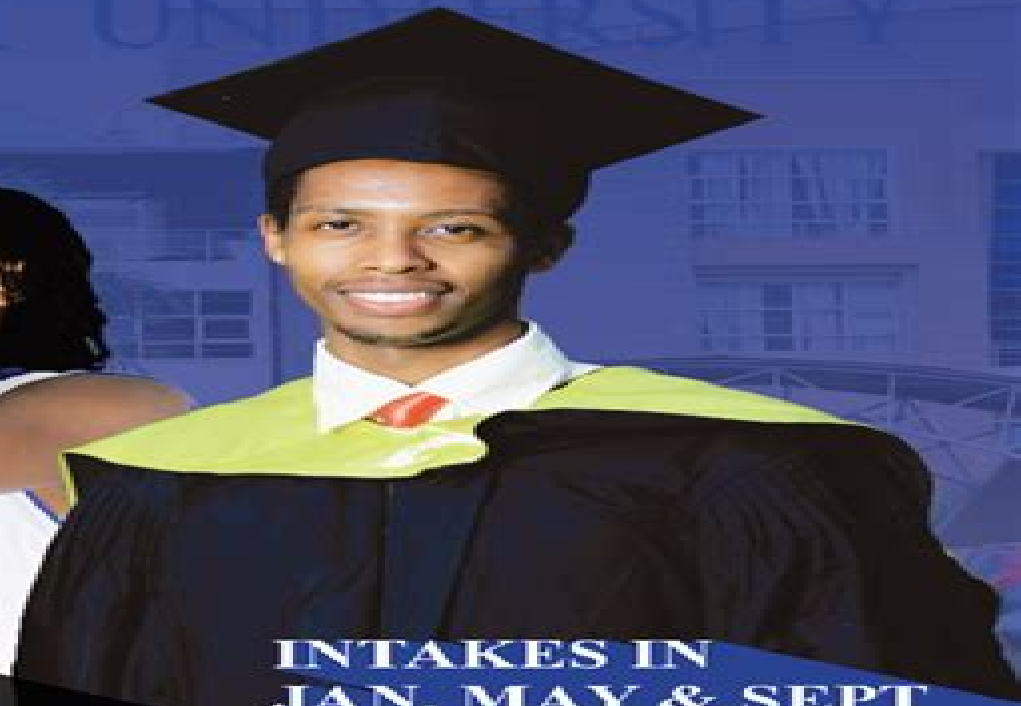




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 what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is
 a helpful way to remember this

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you $23 \text{ Fe} + \text{Co}^{2+} \rightarrow \text{Fe}^{2+} + \text{Co}$
 oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$ reduction $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni}^{2+} \rightarrow 3 \text{ Ag} + \text{Ni}$ oxidation reduction $25 \text{ Cu} + \text{Pb}^{2+} \rightarrow \text{Cu}^{2+} + \text{Pb}$ oxidation reduction 26 O_2

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 magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent
 $\text{a } \text{I}_2 + \text{O}_5 \rightarrow \text{I}_2 + \text{O}_5$
 $5\text{Co} + \text{I}_2 \rightarrow 5\text{Co}^{2+} + \text{I}_2$ $5 \text{ Co} + \text{I}_2 \rightarrow 5 \text{ Co}^{2+} + \text{I}_2$

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 reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation
 state and oxidation number concepts to identify redox reactions and to keep track of electrons

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 reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ $\text{Sr} \rightarrow \text{Sr}^{2+}$ oxidized reducing agent $\text{O}_2 \rightarrow \text{O}^{2-}$ reduced ox ag $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ $\text{Li} \rightarrow \text{Li}^+$ oxidized red ag
 $\text{S} \rightarrow \text{S}^{2-}$ reduced ox ag $3 \text{ Cs} + \text{Br}_2$

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