

Modeled after AP 2013 #21:

1. $2\text{C}_2\text{H}_6(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of $\text{CO}_2(\text{g})$ is produced?
 A. 0.50 mol
 B. 1.50 mol
 C. **4.50 mol**
 D. 9.00 mol
2. $\text{C}_2\text{H}_6(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of reactants consumed when 2.00 mole of $\text{CO}_2(\text{g})$ is produced?
 A. 0.50 mol
 B. 1.00 mol
 C. **3.50 mol**
 D. 7.00 mol
3. $2\text{C}_2\text{H}_6(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of products made when 3.50 mole of $\text{O}_2(\text{g})$ is consumed?
 A. 2.00 mol
 B. 3.00 mol
 C. **5.00 mol**
 D. 10.0 mol
4. $\text{C}_2\text{H}_6(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
In the reaction represented above, what is the total number of moles of products made when 3.50 mole of $\text{O}_2(\text{g})$ is consumed?
 A. 1.00 mol
 B. **2.00 mol**
 C. 4.00 mol
 D. 8.00 mol

Modeled after AP 2014g #48:

3. $4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$
A 7.0 mol sample of $\text{NH}_3(\text{g})$ and a 10. mol sample of $\text{O}_2(\text{g})$ are placed in a previously evacuated container and react according to the equation above. How many moles of $\text{NO}(\text{g})$ are produced when one of the reactants has been totally consumed?
 A. 2.5 mol
 B. **7.0 mol**
 C. 8.0 mol
 D. 28 mol
6. $\text{C}_2\text{H}_6(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{CH}_4(\text{g})$
Samples containing 4.0 mol $\text{C}_2\text{H}_6(\text{g})$ and 5.0 mol $\text{H}_2(\text{g})$ are placed in a previously evacuated container and allowed to react according to the equation above. When one of the reactants is completely consumed, how much product is formed?

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Table of Contents Zumdahl Ap Chemistry Review Questions Answers

1. Understanding the eBook Zumdahl Ap Chemistry Review Questions Answers
 - The Rise of Digital Reading Zumdahl Ap Chemistry Review Questions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Zumdahl Ap Chemistry Review Questions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Zumdahl Ap Chemistry Review Questions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Zumdahl Ap Chemistry Review Questions Answers
 - Personalized Recommendations
 - Zumdahl Ap Chemistry Review Questions Answers User Reviews and Ratings
 - Zumdahl Ap Chemistry Review Questions Answers and Bestseller Lists

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 - Zumdahl Ap Chemistry Review Questions Answers Budget-Friendly Options
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7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Zumdahl Ap Chemistry Review Questions Answers
 - Highlighting and Note-Taking Zumdahl Ap Chemistry Review Questions Answers
 - Interactive Elements Zumdahl Ap Chemistry Review Questions Answers
8. Staying Engaged with Zumdahl Ap Chemistry Review Questions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Zumdahl Ap Chemistry Review Questions Answers
9. Balancing eBooks and Physical Books Zumdahl Ap Chemistry Review Questions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Zumdahl Ap Chemistry Review Questions Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Zumdahl Ap Chemistry Review Questions Answers
 - Setting Reading Goals Zumdahl Ap Chemistry Review Questions Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Zumdahl Ap Chemistry Review Questions Answers
 - Fact-Checking eBook Content of Zumdahl Ap Chemistry Review Questions Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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