

SOLUTIONS MANUAL FOR

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Fundamental  
Mechanics of Fluids  
Fourth Edition

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by

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I.G. Currie

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web a at atmospheric pressure helium can exist in all three phases as well as a supercritical fluid phase near absolute zero choice b solid and gaseous helium never exist in equilibrium with each other at any temperature or pressure b solid and gaseous helium never exist in equilibrium with each other at any temperature or pressure choice c

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web stant is the foundation of thermochemistry it allows us to describe heat flow in terms of the change in a state function enthalpy change is path independent note with regard to choice b that enthalpy change equals internal energy change only if the volume of the system remains constant and thus no thermody namic work is performed 5 b

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web 1 answer a explanation the value of  $\Delta h$  is negative therefore the reaction is exothermic or heat releasing 2 answer a 1 explanation the entire reaction was multiplied by 2 therefore  $\Delta h$  must also be halved 3 answer c explanation option c is incorrect because 566 kj mol is the  $\Delta h$  per mole of  $\text{O}_2$  and not for 2 moles

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web thermochemical equations equations that represent both mass and enthalpy changes  $\text{H}_2\text{O (s)} \rightarrow \text{H}_2\text{O (l)}$   $\Delta h = 6.01 \text{ kJ/mol}$  this is an endothermic process it requires 6.01 kJ to melt one mole of ice  $\text{H}_2\text{O (s)}$  the enthalpy value will change if the number of moles varies from the 1 1 reaction stoichiometry this is an exothermic process

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web in this set of practice questions we will summarize the main concepts of thermochemistry such as the relationship between internal energy work and heat exothermic and endothermic process heat capacity constant pressure calorimetry constant volume calorimetry the enthalpy the standard enthalpies of formation and their use in

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard

enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy  $e$  of a system in any process is equal to the heat  $q$  added to the system plus the work  $w$  done on the system by its surroundings  $e$

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