

THE MOLE: LAB ACTIVITY

Station 1:

A) Determine the mass of one drop of water by measuring the mass of 50 drops of water.

Mass of dish (or beaker)	=	_____	g
Mass of dish (or beaker) + 50 drops water	=	_____	g
Mass of 50 drops water	=	_____	g
Mass of 1 drop of water	=	_____	g

B) How many molecules are there in one drop of water?

Remember 1 mole of a substance is 6.02×10^{23} particles (atoms or molecules).
Therefore, 6.02×10^{23} molecules of water will weigh 18.02 g/mol

- Find # of moles of water for 1 drop $n = (\text{mass})/(\text{Molar Mass})$
- # of molecules in one drop of water $= n_{\text{water}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}}$
 $= ? \text{ molecules}$

Station 2

Calculate the number of copper atoms in a penny.

- Find # of moles of Cu $n = (\text{mass})/(\text{Molar Mass})$
- # of Cu atoms in a penny $= n_{\text{Cu}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}}$
 $= ? \text{ atoms}$

Station 3

A. Calculate the number of moles in one spoon of sugar. Assume that sugar is made up entirely of sucrose molecules ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$).

$$M_{\text{sucrose}} = 342.34 \text{ g/mol}$$

$$n_{\text{sugar}} = \frac{M_{\text{sucrose}}}{M_{\text{sucrose}}}$$

$$= \underline{\hspace{2cm}} \text{ mol}$$

B. Calculate the number of molecules of sugar in one spoon of sugar.

$$\# \text{ of molecules sucrose} = n_{\text{sucrose}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}}$$

$$= \underline{\hspace{2cm}} \text{ mol} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}}$$

Understanding The Mole Lab Activity Answers

Michael Brown



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