

17.	n-hexane	69
18.	Toluene	110

➤ OBSERVATION AND RESULT

The boiling point of the supplied unknown liquid is °C.

➤ PRECAUTIONS

1. Select a suitable liquid medium if the boiling point of the liquid is higher than the boiling point of water.
2. The lower portion of boiling tube should not touch the bottom of beaker.
3. Heating should be done gently.

EXPERIMENT–3.3

➤ AIM OF THE EXPERIMENT

To prepare crystals of potash alum [$K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O$].

➤ MATERIALS REQUIRED

Beaker (250 mL), glass rod, funnel, wash bottle, china dish, filter papers, tripod stand, funnel stand, burner, potassium sulphate (8 g), aluminium sulphate (12 g), dil. sulphuric acid.

➤ THEORY

■ Potash alum is commonly known as 'phitkari'. It is a double salt of potassium sulphate and aluminium sulphate. The molecular formula of potash alum is $K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O$. It can be prepared by heating equimolar quantities of both the salts in water upto crystallisation point followed by cooling.

■ Chemical equation –



➤ PROCEDURE

1. Take 8 g of potassium sulphate in a test tube or beaker, add distilled water drop by drop and prepare saturated solution. Dissolve maximum quantity of salt in minimum quantity of water. Stir it till it becomes clear as shown in [Fig. 3.6].
2. Take 250 mL beaker and add 12 g of aluminium sulphate to it. Add dilute sulphuric acid and stir it by a glass rod. Now, add little water drop by drop and dissolve the salt to prepare saturated solution. If salt does not dissolve then warm the beaker and stir it.
3. Take funnel stand and put a clean china dish under it. Place funnel in the stand with filter papers. Filter both the solutions one by one and collect the filtrate in china dish in [Fig. 3.7].
4. Heat the china dish containing saturated solutions of K_2SO_4 and $Al_2(SO_4)_3$ on a tripod stand containing wire gauze as shown in [Fig. 3.8].

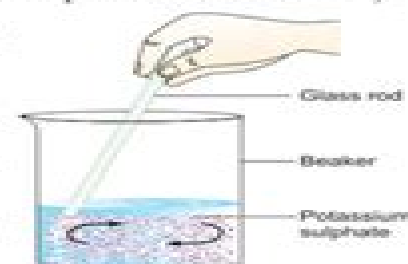


Fig 3.6 Preparation of saturated solution of K_2SO_4

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